

STOICHIOMETRY Worksheet

1. How many moles of sodium will react with water to produce 4.0 mol of hydrogen in the following reaction?



2. Ammonia is made industrially by reacting nitrogen and hydrogen under high pressure and temperature and in the presence of a catalyst. The equation is:



If 4.0 mol of H_2 react, how many moles of NH_3 will be produced?

3. Phosphorous burns in oxygen to produce phosphorous oxide in the following reaction:



a. What amount (in moles) of phosphorous will be needed to produce 3.25 mol of P_4O_{10} ?

b. How many moles of P_4O_{10} are produced from 2.74 moles of phosphorous?

4. How many moles of lithium chloride will be formed by the reaction of chlorine with 0.046 mol of lithium bromide in the following reaction?



5. Aluminum will react with sulfuric acid in the following reaction.



How many moles of H_2SO_4 will react with 18 mol of Al?

Chapter Stoichiometry Worksheet Answers

Sir Tarun Rupani



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cut method A short cut method for solving reacting quantities involving masses and volumes of gases is also given Chapter 3 describes the calculations involved in the practical determination of molecular and empirical formulas A clear meaning of percentage composition of mass is provided and used to solve problems in a step by step manner In chapter 4 we discuss percentage yield and purity A number of examples are given to illustrate how formulas of yield and purity are used in various circumstances A student will find these examples helpful in relating different formulas for percentage purity The last chapter introduces a graphical method for reacting quantities The method may provide a new way of looking at chemical reactions Examples are given to illustrate the method including how it can be used to determine limiting reagents It is hoped that the book will provide all the necessary knowledge and skills to students studying an introductory chemistry course Teachers may also find this book a good resource for their lessons in stoichiometry *Stoichiometry* Sydney Young, 2012-01 Unlike some other reproductions of classic texts 1 We have not used OCR Optical Character Recognition as this leads to bad quality books with introduced typos 2 In books where there are images such as portraits maps sketches etc We have endeavoured to keep the quality of these images so they represent accurately the original artefact Although occasionally there may be certain imperfections with these old texts we feel they deserve to be made available for future generations to enjoy Study Guide for Chemical Principles [by] Steven S. Zumdahl Steven S. Zumdahl, Paul B. Kelter, 1995 The Study Guide reflects the unique problem solving approach taken by the Chemical Principles text The new edition of the Study Guide includes many new worked out examples *You Can Do Chemistry* Sunil Tanna, 2018-11-19 A comprehensive guide to performing mole and stoichiometric calculations with numerous examples as well as questions and answers Covers calculations relating to solids solutions gases and electrolysis plus as limiting and excess reactants chemical yields atom economy and much more Fully up to date with the last international standards including the revised definition of mole which was agreed on November 16th 2018

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