

1. Ion-dipole force always require \_\_\_\_
  - a. an ion and a water molecule
  - b. a cation and a water molecule
  - c. an anion and a polar molecule
  - ✓ d. an ion and a polar molecule
  - e. a polar and a nonpolar molecule
2. Which substance in each of the following pairs would you expect to have a higher boiling point?
  - a.  $O_2$  and  $Cl_2$   **$Cl_2$**
  - b.  $SO_2$  and  $CO_2$   **$SO_2$**
  - c.  $HF$  and  $HI$   **$HF$**
  - d.  $F_2$  and  $LiF$   **$LiF$**
3. The following compound have the same molecular forces ( $C_6H_{10}$ ). Which one would you expect to have a higher boiling point? Explain your answer.



Molecule A has higher boiling point. Van der Waals forces are dependent on surface area. Both molecules have dispersion forces and have the same molecular mass but molecule A has larger surface area thus stronger IMF.

4.  $HCl$  (MW=36.5) has a dipole moment of 1.03 D and a boiling point of 190K.  $HBr$  (MW=80.9) has a dipole of 0.79 D and a boiling point of 206K. Which of the following statements is true?
  - a.  $HBr$  is more polar.
  - b.  $HCl$  has stronger intermolecular forces.
  - c.  $HCl$  has stronger London dispersion forces.
  - d. Both molecules have hydrogen bonding.
  - ✓ e.  $HBr$  has stronger intermolecular forces.
5. Hexane,  $C_6H_{14}$  (MW=86) has a boiling point of 68°C. Ethanol,  $CH_3CH_2OH$  (MW=46) has a boiling point of 78°C. Mark each of the following statements as TRUE or FALSE.
  - a. Ethanol must have stronger intermolecular attraction, based on its higher boiling point. **TRUE**
  - b. Ethanol has a higher boiling point because of greater London dispersion force. **FALSE**
  - c. Both hexane and ethanol have hydrogen bonding. **FALSE**
  - d. Ethanol has a higher boiling point due to hydrogen bonding. **TRUE**
  - e. Hydrogen bonding and London dispersion forces are at cross purposes here. (One favors ethanol, the other favors hexane.) In this case, the hydrogen bonding "wins". **TRUE**

# Chapter Section 12 Liquids And Solids Answer Key

**Martin Stuart Silberberg**



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