CHAPTER 10 The Mole

Section 1 Measuring Matter

period 320-324

Practice Problems

pages 323-324

 Zinc (Zn) is used to form a corrosion-inhibiting surface on galvanized steel. Determine the number of Zn atoms in 2.50 mol of Zn.

 $=1.51\times10^{24}$ atoms of Zn

Calculate the number of molecules in 11.5 mol of waster (Hi-O).

 Silver nitrate (AgNO₁) is used to make several. different silver halides used in photographic films. How many formula units of AgNO3 are there in 3.25 mod AgNO₁?

3.25 mol AgNO,
$$\times \frac{6.02 \times 10^{10} \text{ formula units AgNO,}}{1 \text{ mol AgNO,}}$$

= 1.96×10^{14} formula units AgNO,

4. Challenge Calculate the number of oxygen atoms in 5.00 mol of oxygen molecules. Oxygen is a diatomic molecule, O₂.

5.00 mol
$$O_2 \times \frac{6.02 \times 10^{21} \text{ molecules } O_2}{1 \text{ mol} O_2}$$

- 5. How many moles contain each of the following?
 - a. 5.75×10²⁴ atoms Al

$$5.75 \times 10^{14}$$
 atoms Al $\times \frac{1 \text{ mol Al}}{6.02 \times 10^{10} \text{ atoms Al}}$
= 9.55 mol Al

b. 2:50×10²⁰ atoms Fe.

Chemistry: Matter and Change

- 6. Challenge Identify the representative particle for each formula and convert the given number of representative particles to moles.
 - a. 3.75 × 10²⁴ CO₂

The representative particle is a molecule.

3.75×1014 molecules CO,

× 1 mol CO; 6.02×10²³ molecules CO; = 6.23 mol CO;

b. 3.58×10²³ ZnCb

The representative particle is a formula unit.

3.58×1023 formula units ZnCl,

× 1 mol ZnCl; 6.02×10²³ formula units ZnCl,

0.595 mol ZnCl.

Section 1 Review

eserge 324

Explain why chemists use the mole.

Chemists use the mole because it is a convenient way of knowing how many representative particles are in a sample.

8. State the mathematical relationship between Avogadro's number and 1 mol.

One mole contains Avogadro's number (6.02×10²³) of representative particles.

Chapter The Mole Study Guide Answers

Robinson

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