

Chapter 5: Thermochemistry

Problem Set with Answers

1. A system receives 425 J of heat and delivers 425 J of work to its surroundings. What is the change in internal energy of the system (in J)?

Answer: 0 J

2. Complete combustion of 1.0 metric ton of coal (assuming pure carbon) to gaseous carbon dioxide releases 3.3×10^{10} J of heat. Convert this energy to (a) kilojoules; (b) kilocalories

Answer: (a) 3.3×10^7 kJ (b) 7.9×10^6 kcal

3. Classify the following processes as exothermic or endothermic: (a) freezing of water; (b) boiling of water; (c) digestion of food; (d) a person running; (e) a person growing; (f) wood being chopped; (g) heating with a furnace.

Answer: (a) exothermic (b) endothermic (c) exothermic (d) exothermic
(e) endothermic (f) endothermic (g) exothermic

4. Calculate q when 12.0 g of water is heated from 20 °C to 100 °C.

Answer: 4.0×10^3 J

5. A 295-g aluminum engine part at an initial temperature of 3.00°C absorbs 85.0 kJ of heat. What is the final temperature of the part (c of Al = 0.900 J/g·K)?

Answer: 323°C

6. Ethylene (C_2H_4) is the starting material for the preparation of polyethylene. Although typically made during the processing of petroleum, ethylene occurs naturally as a fruit-ripening hormone and as a component of natural gas.
(a) The heat of reaction for the combustion of C_2H_4 is -1411 kJ/mole. Write a balanced thermochemical equation for the combustion of C_2H_4 .
(b) How many grams of C_2H_4 must burn to give 70.0 kJ of heat?

Answer: (a) $\text{C}_2\text{H}_4 (\text{g}) + 3\text{O}_2 (\text{g}) \rightarrow 2\text{CO}_2 (\text{g}) + 2\text{H}_2\text{O} (\text{l})$; $\Delta H_{\text{rxn}} = -1411 \text{ kJ}$
(b) 1.39 g C_2H_4

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RD Boyd



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