

## Chapter Two: Principles of Geochemistry and Thermodynamics

### ❖ Ionic radii and Ionic substitution in crystals

- ❖ A crystal is any substance whose atoms are arranged in a regular, periodically repeated pattern.
- ❖ Ionic radius is a measure of the size of an atom's ion in a crystal lattice. Ionic radius is measured in either picometres (pm) or Angstrom ( $\text{\AA}$ ), with  $1 \text{ \AA} = 100 \text{ pm}$ .
- ❖ Most minerals exhibit a considerable range in chemical composition. Such variation results from the replacement of one ion or ionic group by another in a particular structure.
- ❖ This phenomenon is termed ionic substitution, or solid solution. The degree of substitution may be influenced by various factors, with the size of the ion (Ionic radius) being the most important.
- Ions of two different elements can freely replace one another only if their ionic radii differ by approximately 15 percent or less.
- Limited substitution can occur if the radii differ by 15 to 30 percent, and a difference of more than 30 percent makes substitution unlikely.
- ❖ The temperature at which crystals grow also plays a significant role in determining the extent of ionic substitution. The higher the temperature, the more extensive is the thermal disorder in the crystal structure and the less exacting are the spatial requirements.
- As a result, ionic substitution that could not have occurred in crystals grown at low temperatures may be present in those

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