

Chapter 3:

Section 3.1

1. Electronegativity is the affinity of an atom for electrons of a covalent bond.
In a water molecule, because the Oxygen atom has greater electronegativity, it attracts the shared pair of electrons with Hydrogen towards itself, becoming partially negatively charged. This causes the O-H bond to become polar and attract adjacent O-H bonds to create strong intermolecular Hydrogen bonds.
2. The partial + charge of each H and partial - charge of each O would repel each other.
3. The equal electronegativity of both atoms would make the O-H bond non-polar and prevent formation of Hydrogen bonds which would affect the properties of water such as its surface tension, adhesion/cohesion, and boiling point.

Section 3.2

1. The cohesion and adhesion of water molecules both toward themselves and towards the cells of the tree (due to the polarity of water) leads it to have strong intermolecular forces (H-bonds) that draw the molecules together even against gravity.
2. The humidity of the air prevents evaporation of sweat off the skin, preventing cooling of the skin.
3. When water freezes, it expands. Therefore, if it is present within the cracks of a boulder, it can split the boulder apart.
4. Concentration is moles/L. So for 1L of blood, there would be $(1.3 \times 10^{-10}) \times$ Avogadro's number (6.26×10^{26}) molecules of ghrelin.
5. The water strider's hydrophobic legs help it repel the water and walk on its surface. If it was hydrophilic, it would pull water drops up its legs and possibly weigh it down, making it more difficult to walk.

Chapter 5

Section 5.1

1. The four main classes of macromolecules are:
Carbohydrates
Proteins
Nucleic acids (DNA, RNA)
Lipids (fats)

Lipids are not polymers

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Martha R. Taylor



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