

Name: _____ Date: _____ Period: _____

Chapter 11 Study Guide

(due Wednesday)

- 1.) Electrons are found in the electron clouds of an atom.
- 2.) All atoms of the same element contain the same number of protons.
- 3.) Electrons have no electric charge.
- 4.) The mass of an element is the number of protons and neutrons in the nucleus.
- 5.) The atomic mass is an average of the masses of all naturally occurring isotopes of an element.
- 6.) How many **protons** does an atom with an **atomic number of 23** and mass number of 51 have?
Protons: 23
- 7.) What determines the **identity of an element**? **Atomic #**
- 8.) **Isotopes** exist because atoms of the same element can have **different numbers of** neutrons.
- 9.) What did Democritis, Dalton, Thomson, Rutherford, and Bohr all have in **common**?
They all contributed ideas for the atomic theory

-
-
- 10.) An atom of gold with **79 protons**, **79 electrons** and **118 neutrons** would have a mass number of 197 amu.
 - 11.) The **smallest particle** into which an element can be divided and still be the same substance is a(n) atom.
 - 12.) **Mass number** includes protons and neutrons.
 - 13.) Which part of an atom has the **least mass**? electrons
 - 14.) Name the **charges** for the following:

Protons: + positive Neutrons: no charge Electrons: - negative

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