

## Chapter 8 Balancing Equations Set # 3

Using the equations from Practice set #1 and #2 perform the following:

1. Balance the equation
2. Classify the balanced equation using the following choices:
  - Synthesis
  - Decomposition
  - Single Displacement (Cation or Anion)
  - Combustion (Hydrocarbon or Carbohydrate)
  - Double Displacement / Double Displacement (Acid/Base)
  - Other

### Set #1

1.  $\text{HF (g)} \longrightarrow \text{H}_2 \text{ (g)} + \text{F}_2 \text{ (g)}$  \_\_\_\_\_
2.  $\text{Al (s)} + \text{I}_2 \text{ (s)} \longrightarrow \text{AlI}_3 \text{ (aq)}$  \_\_\_\_\_
3.  $\text{FeO (s)} + \text{O}_2 \text{ (g)} \longrightarrow \text{Fe}_2\text{O}_3 \text{ (s)}$  \_\_\_\_\_
4.  $\text{C}_4\text{H}_{10} \text{ (s)} + \text{O}_2 \text{ (g)} \longrightarrow \text{CO}_2 \text{ (g)} + \text{H}_2\text{O (l)}$  \_\_\_\_\_
5.  $\text{Na (s)} + \text{MgI}_2 \text{ (aq)} \longrightarrow \text{NaI (aq)} + \text{Mg (s)}$  \_\_\_\_\_
6.  $\text{AgNO}_3 \text{ (s)} + \text{Na}_2\text{S (aq)} \longrightarrow \text{Ag}_2\text{S (s)} + \text{NaNO}_3 \text{ (aq)}$  \_\_\_\_\_
7.  $\text{K}_2\text{SO}_4 \text{ (aq)} + \text{Mg(NO}_3)_2 \text{ (aq)} \longrightarrow \text{MgSO}_4 \text{ (ppt)} + \text{KNO}_3 \text{ (aq)}$  \_\_\_\_\_
8.  $\text{Al}_2(\text{CO}_3)_3 \text{ (s)} \longrightarrow \text{Al}_2\text{O}_3 \text{ (s)} + \text{CO}_2 \text{ (aq)}$  \_\_\_\_\_
9.  $\text{Li (s)} + \text{AuCl}_3 \text{ (aq)} \longrightarrow \text{LiCl (aq)} + \text{Au (s)}$  \_\_\_\_\_
10.  $\text{Li}_2\text{CrO}_4 \text{ (aq)} + \text{BaCl}_2 \text{ (aq)} \longrightarrow \text{LiCl (aq)} + \text{BaCrO}_4 \text{ (ppt)}$  \_\_\_\_\_
11.  $\text{C}_{12}\text{H}_{22}\text{O}_{11} \text{ (s)} + \text{O}_2 \text{ (g)} \longrightarrow \text{CO}_2 \text{ (g)} + \text{H}_2\text{O (l)}$  \_\_\_\_\_
12.  $\text{Ca(OH)}_2 \text{ (aq)} + \text{H}_2\text{SO}_4 \text{ (aq)} \longrightarrow \text{CaSO}_4 \text{ (aq)} + \text{H}_2\text{O (l)}$  \_\_\_\_\_

# Chapter 8 Balancing Equations Prentice Hall

**Stephen F. Davis**



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