

Chemistry Student Edition - Basic Answer Key

Chapter 20: Entropy and Free Energy

Entropy

Questions

1. Define entropy.
2. State the second law of thermodynamics.
3. For each of the following situations, state whether there is an increase or decrease in entropy. Explain your reasoning in each case.
 - a) liquid water freezes.
 - b) a car is in a collision that completely demolishes it.
 - c) wood burns.
 - d) $\text{I}_{2(s)} \rightarrow \text{I}_{2(g)}$.
 - e) $2 \text{Mg}_{(s)} + \text{O}_{2(g)} \rightarrow 2 \text{MgO}_{(s)}$.

Use the entropy values from <http://chemed.chem.wisc.edu/chempaths/Table-of-Standard-Molar-Entropies-1184.html> for the following calculations:

4. Calculate $\Delta S^{\circ}_{\text{rxn}}$ for the reaction $\text{H}_{2(g)} + \text{Cl}_{2(g)} \rightarrow 2\text{HCl}_{(g)}$.
5. Calculate $\Delta S^{\circ}_{\text{rxn}}$ for reactions d and e of question 3, and comment on how well your predictions matched the actual entropy changes.
6. Which form of carbon has a higher degree of organization: diamond or graphite? Explain your answer.
7. Compare the S° of ethane with that of ethanol (159.9 J/K·mol). Explain why ethane has a higher standard entropy value than ethanol. (Hint: the boiling point of ethane is -89°C , and the boiling point of ethanol is 78°C . At what temperature are standard entropy values tabulated?)

Answers

1. The degree of randomness in a system
2. Entropy increases in a spontaneous process
3.
 - a) Decrease, more order in a solid than in a liquid
 - b) Increase, more disorder in the wreck
 - c) Increase, more disorder in two gases than in one gas and one solid
 - d) Increase, more disorder in a gas than in a solid
 - e) $2 \text{Mg}_{(s)} + \text{O}_{2(g)} \rightarrow 2 \text{MgO}_{(s)}$.
4.
$$\Delta S^{\circ}_{\text{rxn}} = \sum n_p S^{\circ} \text{products} - \sum m_r S^{\circ} \text{reactants}$$
$$= 2(186.9) - [130.7 + 223.1]$$
$$= 20 \text{ J/mol}$$
5. d.
$$\Delta S^{\circ}_{\text{rxn}} = \sum n_p S^{\circ} \text{products} - \sum m_r S^{\circ} \text{reactants}$$
$$\Delta S^{\circ}_{\text{rxn}} = 260.7 - 116.1$$

Chapter Chemistry Answer Key

Padhraic Smyth



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