



# COVALENT BONDING

## Practice Problems

Do your answers to these very practice problems.

### SECTION 8.1 MOLECULAR COMPOUNDS

1. Classify each of the following compounds as ionic or covalent.

- a.  $\text{Na}_2\text{O}$  ionic      b.  $\text{CH}_4$  covalent      c.  $\text{HCl}$  covalent  
d.  $\text{KNO}_3$  ionic      e.  $\text{H}_2\text{O}$  covalent

2. Write the formula for the following molecular compounds.

- a. carbon dioxide      b. sulfur dioxide      c. carbon monoxide  
d. nitrogen      e. sulfur      f. phosphorus

3. Write the formula for the following molecular compounds: nitrogen.

4. Write the formula for the following molecular compounds: carbon dioxide, carbon monoxide, sulfur dioxide, sulfur trioxide, dinitrogen pentoxide, and phosphorus pentoxide.

5. Write the formula for the following molecular compounds: carbon tetrachloride, carbon disulfide, carbon monoxide, carbon dioxide, and carbon trioxide.

### SECTION 8.2 THE NATURE OF COVALENT BONDING

1. Draw the Lewis structure for hydrogen fluoride,  $\text{HF}$ .



2. Draw the Lewis structure for methane,  $\text{CH}_4$ .



3. Draw the Lewis structure for carbon dioxide,  $\text{CO}_2$ .



4. Draw the Lewis structure for acetylene,  $\text{C}_2\text{H}_2$ .

5. Draw the Lewis structure for the following:  $\text{H}_2\text{O}$ ,  $\text{NH}_3$ ,  $\text{CH}_4$ ,  $\text{CO}_2$ ,  $\text{C}_2\text{H}_2$ ,  $\text{C}_2\text{H}_4$ ,  $\text{C}_2\text{H}_6$ ,  $\text{C}_3\text{H}_8$ ,  $\text{C}_4\text{H}_{10}$ ,  $\text{C}_5\text{H}_{12}$ .

Answer key:



### SECTION 8.3 BONDING: THE ORBITAL

1. Write the formula for the bond between the carbon atoms in acetylene,  $\text{C}_2\text{H}_2$ . The bond between the two carbon atoms.

2. Write the formula for the bond between the two carbon atoms in acetylene,  $\text{C}_2\text{H}_2$ . The bond between the two carbon atoms.

# Chapter 8 Covalent Bonding Practice Problems Answer Key

**Jessica J Manson**



## **Chapter 8 Covalent Bonding Practice Problems Answer Key:**

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