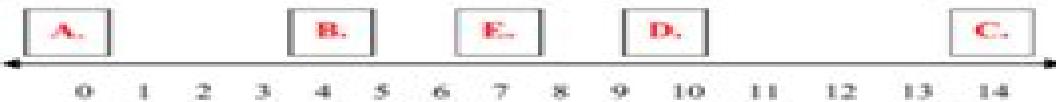


Unit 12: Chapter 8 – Solutions, Acids, and Bases Study Guide Answer Key

1. Define:
 - a. Solute - a substance whose particles are dissolved in a solution.
 - b. Solvent - a substance in which a solute dissolves.
 - c. Solution - a mixture that forms when substances dissolve and form a homogeneous mixture.
 - d. Identify and explain what each part is for lemonade. solvent - water; solutes - lemon juice & sugar, solution - lemonade
 - e. Saturated solution - a solution that contains as much solute as the solvent can normally hold at a given temp.
 - f. Unsaturated solution - a solution in which more solute can be dissolved at a given temperature.
 - g. Supersaturated solution - a solution that contains more solute than the solvent can normally hold at a given temperature.
2. List and describe three (3) ways we can increase the rate of dissolving of a solid.
 - a. Increasing temperature - heating
 - b. Increasing surface area - smaller particles by crushing
 - c. Stirring/shaking
3. List and describe three (3) ways we can increase the rate of dissolving of a gas.
 - a. Decreasing temperature - cooling
 - b. Increasing pressure
 - c. Not stirring/shaking
4. Does salt dissolve faster in hot water or in cold water? Why?
Salt dissolves faster in hot water because salt is a solid and solid solutes dissolve faster when heated.
5. Draw a pH scale and label it with the following :

a. strong acids
b. weak acids
c. strong bases
d. weak bases
e. neutral



0 1 2 3 4 5 6 7 8 9 10 11 12 13 14
6. Name five (5) properties of an acid.
pH less than 7; sour taste; corrosive; electrolytes; turn litmus red; react with metals to form H₂ gas
7. Give 3 examples of acidic items: lemon juice, car battery acid, gastric juice, vinegar
8. Name five (5) properties of a base.
pH greater than 7; bitter taste; corrosive; electrolytes; turn litmus blue; feel slippery
9. Give 3 examples of basic items: antacids/laxatives, detergents/soaps, bleach
10. An acid is defined as a compound that produces H₃O⁺ or is a proton donor.
11. A base is defined as a compound that produces OH⁻ or is a proton acceptor.

Label the following as strong acids (SA), weak acids (WA), strong bases (SB), weak bases (WB), or neutral (N).



- | | |
|---|--------------------------------|
| 12. Battery acid (hydrofluoric acid) pH=1 SA | 16. Baking soda pH=8 WB |
| 13. Lye pH=13 SB | 17. Lemon juice pH=2 SA |
| 14. Orange juice pH=3 WA | 18. Lemon soap pH=6 WA |
| 15. Soap pH=9 WB | 19. Pure water pH=7 N |

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Lei Shi



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