

CHAPTER 8: ACIDS & BASES
STANDARD SL-L & HL-M
STUDY GUIDE

SECTION 8:1

1. What are the major differences between Arrhenius, Bronstead-Lowery and Lewis definition of acid and bases

Arrhenius defines acids as a species that increases the concentration of hydrogen ions in solution and a bases is a species that increases the concentration of hydroxide ions in solution.

Bronstead-Lowry defines acids as a species that is a proton (H^+) donor and a base as a proton (H^+) acceptor.

Lewis defines acids as an electron pair acceptor and a base as an electron pair donor

2. What does amphoteric (amphiprotic) mean? Something that can act as an acid or base

3. What enables certain species to be amphoteric? They must have the ability to accept a H^+ or donate a H^+ .

4. Give an example of how water can act as an acid.



5. Give an example of how water can act as a base.



6. Explain what a dative bond is and how it relates to Lewis acids and bases?

A dative bond is a bond that forms when both electrons in the pair comes from the same molecule. A Lewis bases contribute the electrons with Lewis acids thus creating a datative bond.

7. What is a ligand and how does it relate to Lewis acids and bases?

A ligand is a Lewis base that donates a lone pair and usually surrounds Lewis acids in a fixed ratio.

8. What is a complex ion and how does it relate to Lewis acids and bases? A complex ion forms when a ligand surrounds a metal ion (usually a transition metal).

9. What role do transition metals play in Lewis acids and bases? Transition metals can be Lewis acids when surrounded by ligands that are Lewis bases

10. Do Book Exercises #1-6 Pgs 148-150 (see answers in back of book)

Exercise #1 Remember that a conjugate acid of a base will be what a base becomes when it gains a H^+ (accepts a proton) That is why SO_3^{2-} (base) becomes HSO_3^- (conjugate acid)

Exercise #2 Remember that a conjugate base of an acid is what an acid becomes when it loses a H^+ (donates a proton) That is why H_2PO_4 (acid) becomes $H_2PO_4^-$ (conjugate base)

Exercise #3 Remember acids donate a H^+ and bases accept H^+ .

Exercise #4 Remember that Lewis bases have unshared pairs of electron and therefore act as electron donors and Lewis acids are often transition metal ions or molecules without unshared pair of electrons

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Table of Contents Chapter Acids And Bases Study Guide

1. Understanding the eBook Chapter Acids And Bases Study Guide
 - The Rise of Digital Reading Chapter Acids And Bases Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter Acids And Bases Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter Acids And Bases Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter Acids And Bases Study Guide
 - Personalized Recommendations
 - Chapter Acids And Bases Study Guide User Reviews and Ratings
 - Chapter Acids And Bases Study Guide and Bestseller Lists
5. Accessing Chapter Acids And Bases Study Guide Free and Paid eBooks

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 - Chapter Acids And Bases Study Guide Enhanced eBook Features
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 - Adjustable Fonts and Text Sizes of Chapter Acids And Bases Study Guide
 - Highlighting and Note-Taking Chapter Acids And Bases Study Guide
 - Interactive Elements Chapter Acids And Bases Study Guide
- 8. Staying Engaged with Chapter Acids And Bases Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter Acids And Bases Study Guide
- 9. Balancing eBooks and Physical Books Chapter Acids And Bases Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter Acids And Bases Study Guide
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter Acids And Bases Study Guide
 - Setting Reading Goals Chapter Acids And Bases Study Guide
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter Acids And Bases Study Guide
 - Fact-Checking eBook Content of Chapter Acids And Bases Study Guide
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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