

Introduction to Acid, Base, and Salt

1. The nature of calcium phosphate present in tooth enamel is
a. Basic b. Amphoteric c. Acidic d. Neutral
Answer ____
2. Which of the following salts has no water of crystallization?
a. Blue vitriol b. Washing soda c. Baking soda d. Gypsum
Answer ____
3. The pH of an HCl solution is 4. What is the molarity of the solution?
a. 4.0 M b. 0.4 M c. 0.0001 M d. 0.001 M
Answer ____
4. Which of the following does not form an acidic salt?
a. Phosphoric acid b. Carbonic acid
c. Hydrochloric acid d. Sulfuric acid
Answer ____
5. The gas that is released when an acid reacts with a metal carbonate is
a. Carbon dioxide b. Hydrogen c. Oxygen d. Nitrogen
Answer ____
6. Two fruits that are rich in citric acid are
a. Lemon and tamarind b. Lemon and orange
c. Orange and tamarind d. Tomato and tamarind
Answer ____
7. The acid present in stinging hair of nettle leaves that causes burning pain
a. Citric acid b. Acetic acid c. Tartaric acid d. Methanoic acid
Answer ____
8. As the pH value of a solution decreases,
a. Number of OH^- ions increases b. Number of H^+ ions increases
c. Number of H^+ ions decreases d. Equal number of H^+ ions and OH^- ions
Answer ____

Chapter Acids Bases And Salts Worksheet Answers

Arshad Iqbal



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pollutants climatology global warming meteorology ozone depletion and troposphere *Acids, Bases, and Salts* Brian J. Knapp, 1998 **Acids, Bases, and Salts** Adams, This note is part of Quality testing 1947, □□□□□□□□ Acids and Bases Brian G. Cox, 2013-01-31

Acids and bases are ubiquitous in chemistry Our understanding of them however is dominated by their behaviour in water Transfer to non aqueous solvents leads to profound changes in acid base strengths and to the rates and equilibria of many processes for example synthetic reactions involving acids bases and nucleophiles isolation of pharmaceutical actives through salt formation formation of zwitter ions in amino acids and chromatographic separation of substrates This book seeks to enhance our understanding of acids and bases by reviewing and analysing their behaviour in non aqueous solvents The behaviour is related where possible to that in water but correlations and contrasts between solvents are also presented Fundamental background material is provided in the initial chapters quantitative aspects of acid base equilibria including definitions and relationships between solution pH and species distribution the influence of molecular structure on acid strengths and acidity in aqueous solution Solvent properties are reviewed along with the magnitude of the interaction energies of solvent molecules with especially ions the ability of solvents to participate in hydrogen bonding and to accept or donate electron pairs is seen to be crucial Experimental methods for determining dissociation constants are described in detail In the remaining chapters dissociation constants of a wide range of acids in three distinct classes of solvents are discussed protic solvents such as alcohols which are strong hydrogen bond donors basic polar aprotic solvents such as dimethylformamide and low basicity and low polarity solvents such as acetonitrile and tetrahydrofuran Dissociation constants of individual acids vary over more than 20 orders of magnitude among the solvents and there is a strong differentiation between the response of neutral and charged acids to solvent change Ion pairing and hydrogen bonding equilibria such as between phenol and phenoxide ions play an increasingly important role as the solvent polarity decreases and their influence on acid base equilibria and salt formation is described

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