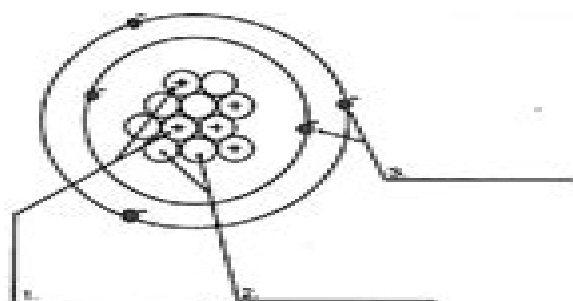


## Atomic Structure Worksheet

Label the parts of an atom on the diagram below.



4. What type of charge does a proton have? +
5. What type of charge does a neutron have? neutral
6. What type of charge does an electron have? -
7. Which two subatomic particles are located in the nucleus of an atom? p + n
8. What is the charge of an atom? neutral
9. What is the charge of an ion? (Explain your answer fully.)  
+ or - gains or loses e<sup>-</sup>
10. If an atom has 35 protons in the nucleus, how many electrons will it have orbiting the nucleus? 35
11. What is the atomic number of the atom in the diagram above? 5

12. What is the atomic mass/mass number of the atom in the diagram above? 11

13. How many protons are in the nucleus of an atom with an atomic number of 15? 15

14. How many electrons are in the nucleus of an atom with an atomic number of 20? 20

15. How many neutrons are in the nucleus of an atom with an atomic number of 25? (use Periodic Table for mass)

16. What is the mass number of an atom with 3 protons, 4 neutrons, and 3 electrons? 7

17. How many neutrons are in the nucleus of an atom that has an atomic mass of 36 and an atomic number of 25? 11

18. The atomic number tells you the number of protons in one atom of an element. It also tells you the number of electrons in a neutral atom of that element. The atomic number gives the "identity" of an element as well as it's location on the Periodic Table. No two different elements will have the same atomic number.

19. The mass number is used to calculate the number of p, n, e? in one atom of an element. In order to calculate the number of neutrons you must subtract the # p from the mass #.  
(atomic #)

# Chapter Atomic Structure Worksheet Answers Pearson

**John E. Parnell**



## **Chapter Atomic Structure Worksheet Answers Pearson:**

**ATOMIC STRUCTURE High School Chemistry Curriculum** Bud Morgan, 2025-07-09 Welcome to our comprehensive guide to Atomic Structure designed specifically for high school students studying chemistry In this book we will explore the fascinating world of atoms the basic building blocks of matter and delve into the complexities of their structure Chapter 1 provides a solid foundation by introducing the concept of atoms and their historical significance in the development of modern chemistry We will examine the key scientists who contributed to our understanding of atomic structure from Democritus to Thomson Rutherford and Bohr Moving on to Chapter 2 we will explore the structure of the atom in more detail including the subatomic particles that make up an atom protons neutrons and electrons We will also discuss the concept of atomic number atomic mass and isotopes In Chapter 3 we will delve into the arrangement of electrons within an atom s electron cloud including energy levels sublevels and orbital shapes We will also discuss the principles of electron configuration and how it impacts an atom s chemical properties Chapter 4 will focus on the periodic table and how it organizes elements based on their atomic structure We will explore periodic trends such as atomic radius ionization energy and electron affinity and discuss how these trends are related to an element s position on the table In our final chapter we will discuss the significance of atomic structure in various chemical reactions including bonding stoichiometry and the behavior of gases We will also explore the importance of understanding atomic structure in fields such as biochemistry environmental science and materials science Throughout this book we will provide clear explanations helpful illustrations and interactive activities to reinforce key concepts Whether you are a visual learner a hands on learner or a theoretical learner this book is designed to meet your needs We hope that this book will not only enhance your understanding of atomic structure but also inspire you to further explore the wonders of chemistry So let s dive in and unlock the mysteries of the atom together

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
















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