

# PRACTICE - TEST on Atomic Structure

Name: \_\_\_\_\_

- 1) Complete the table below – *Note that the particles are not in the same order.* 3 points

	Charge	Location	AMU
Neutron	0	In Nucleus	1
Electron	1-	Electron Cloud	0
Proton	1+	In Nucleus	1

- 2) Complete the table below – 5 points

**Protons + Neutrons = Mass #**

Isotope Symbol	Element Name	Atomic Number	Number of Protons	Number of Neutrons	Mass Number	Number of Electrons
$^{15}\text{N}$	Nitrogen	7	7	8	15	7
$^{28}\text{Al}$	Aluminum	13	13	15	28	13
$^{41}\text{Ca}$	Calcium	20	20	21	41	20
$^{74}\text{Br}$	Bromine	35	35	39	74	35

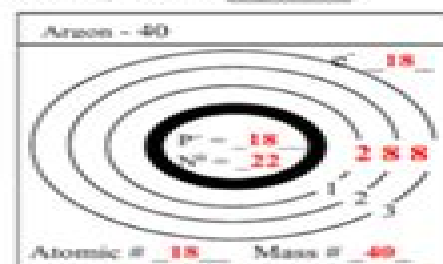
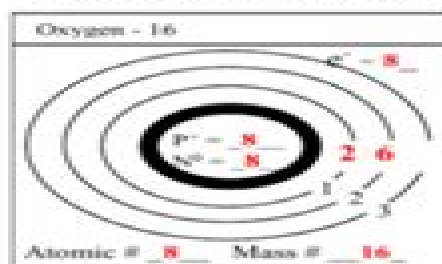
- 3) Complete the table below – 5 points

		Atomic #	P <sup>+</sup>	e <sup>-</sup>	E1	E2	E3	E4	Ve <sup>-</sup>	Group
F	N-m	9	9	9	2	7			7	7A - Halogens
P	N-m	15	15	15	2	8	5		5	5A -
Mg	M	12	12	12	2	8	2		2	2A - Alkali Earth Metals

- 4) Draw the Lewis Structures for the following neutral atoms. 6 points



- 5) Complete the boxes below. Be sure to include the electron configuration on the correct energy levels. Draw the Lewis Structure of each resulting atom. 7 points



# Chapter Atomic Structure Test A Answers

**Peter P. Berlow**



## **Chapter Atomic Structure Test A Answers:**

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