

Chapter 2

Electrochemistry

Intext Questions

Question 2.1:

How would you determine the standard electrode potential of the system $\text{Mg}^{2+} | \text{Mg}$?

Answer:

The standard electrode potential of $\text{Mg}^{2+} | \text{Mg}$ can be measured with respect to the standard hydrogen electrode, represented by $\text{Pt}(s), \text{H}_2(g, 1 \text{ atm}) | \text{H}^+(aq, 1 \text{ M})$.

A cell, consisting of $\text{Mg} | \text{MgSO}_4 (aq, 1 \text{ M})$ as the anode and the standard hydrogen electrode as the cathode, is set up.



Then, the emf of the cell is measured and this measured emf is the standard electrode potential of the magnesium electrode.

$$E^\circ = E^\circ_{\text{R}} - E^\circ_{\text{L}}$$

Here, E°_{R} for the standard hydrogen electrode is zero.

$\therefore$

$$E^\circ = 0 - E^\circ_{\text{L}}$$

$$= -E^\circ_{\text{L}}$$

Question 2.2:

Can you store copper sulphate solutions in a zinc pot?

Answer:

Zinc is more reactive than copper. Therefore, zinc can displace copper from its salt solution. If copper sulphate solution is stored in a zinc pot, then zinc will displace copper from the copper sulphate solution.



Chapter Electrochemistry Answers

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