

## CHAPTER 15 REVIEW

***Acid-Base Titration and pH*****SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

1. Calculate the following values without using a calculator.

- \_\_\_\_\_ a. The  $[\text{H}_3\text{O}^+]$  is  $1 \times 10^{-6}$  M in a solution. Calculate the  $[\text{OH}^-]$ .
- \_\_\_\_\_ b. The  $[\text{H}_3\text{O}^+]$  is  $1 \times 10^{-9}$  M in a solution. Calculate the  $[\text{OH}^-]$ .
- \_\_\_\_\_ c. The  $[\text{OH}^-]$  is  $1 \times 10^{-12}$  M in a solution. Calculate the  $[\text{H}_3\text{O}^+]$ .
- \_\_\_\_\_ d. The  $[\text{OH}^-]$  in part c is reduced by half, to  $0.5 \times 10^{-12}$  M. Calculate the  $[\text{H}_3\text{O}^+]$ .
- \_\_\_\_\_ e. The  $[\text{H}_3\text{O}^+]$  and  $[\text{OH}^-]$  are \_\_\_\_\_ (directly, inversely, or not) proportional in any system involving water.

2. Calculate the following values without using a calculator.

- \_\_\_\_\_ a. The pH of a solution is 2.0. Calculate the pOH.
- \_\_\_\_\_ b. The pOH of a solution is 4.73. Calculate the pH.
- \_\_\_\_\_ c. The  $[\text{H}_3\text{O}^+]$  in a solution is  $1 \times 10^{-3}$  M. Calculate the pH.
- \_\_\_\_\_ d. The pOH of a solution is 5.0. Calculate the  $[\text{OH}^-]$ .
- \_\_\_\_\_ e. The pH of a solution is 1.0. Calculate the  $[\text{OH}^-]$ .

3. Calculate the following values.

- \_\_\_\_\_ a. The  $[\text{H}_3\text{O}^+]$  is  $2.34 \times 10^{-5}$  M in a solution. Calculate the pH.
- \_\_\_\_\_ b. The pOH of a solution is 3.5. Calculate the  $[\text{OH}^-]$ .
- \_\_\_\_\_ c. The  $[\text{H}_3\text{O}^+]$  is  $4.6 \times 10^{-8}$  M in a solution. Calculate the  $[\text{OH}^-]$ .

**PROBLEMS** Write the answer on the line to the left. Show all your work in the space provided.4.  $[\text{H}_3\text{O}^+]$  in an aqueous solution =  $2.3 \times 10^{-3}$  M.

- \_\_\_\_\_ a. Calculate  $[\text{OH}^-]$  in this solution.

# Chapter Review Acid Base Titration Ph Mixed Answers

**Jerry A. Bell, American Chemical  
Society**



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the curriculum of all the medical sciences The present book is an attempt in this direction in the form of a student friendly yet comprehensive and up to date text

**General College Chemistry** Charles William Keenan, Donald C. Kleinfelter, Jesse Hermon Wood, 1980

**New Information Technology in Education** David Hawkrige, 2022-06-01 First published in 1983 New Information Technology in Education surveyed developments in the field of information technology and demonstrated how it could be used to improve the quality of education The book considered the experience of a wide range of countries including the United States Japan and those in Europe While explaining the potential improvements that the new technology could bring this book also reviewed the problem areas and helped educationalists to evaluate the relevance of the new technology for their own work In an age of teaching via Zoom videos it is interesting to take a look at a time when information technology in education was at its nascent stage This book will be of interest to teachers and students of history education technology and pedagogy

**Holt Chemistry** Ralph Thomas Myers, 2004 Introduction to Drug Analysis Robert Francis Raffauf, Victor D. Warner, 1978

*Cehmistry Textbook for College and University USA* Ibrahim Sikder, 2023-06-04 Cehmistry Textbook USA Chemistry John McMurry, 1998

**Merrill Chemistry** Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 Textbook of Biochemistry Thomas M. Devlin, 1986-04-07 An autobiography of a boy who at the age of three fled civil war in Ethiopia by walking with his mother and brother to a Sudanese refugee camp and later moved to Chicago and earned a scholarship to Harvard University Includes recipes and discussion questions

*Technical Education and Industrial Training*, 1964 *Mellor's Modern Inorganic Chemistry* Joseph William Mellor, 1967

*Aqueous Acid-base Equilibria and Titrations* Robert De Levie, 1999 This book will give students a thorough grounding in pH and associated equilibria material absolutely fundamental to the understanding of many aspects of chemistry It is in addition a fresh and modern approach to a topic all too often taught in an out moded way This book uses new theoretical developments which have led to more generalized approaches to equilibrium problems these approaches are often simpler than the approximations which they replace Acid base problems are readily addressed in terms of the proton condition a convenient amalgam of the mass and charge constraints of the chemical system considered The graphical approach of Bjerrum Hagg and Sillen is used to illustrate the orders of magnitude of the concentrations of the various species involved in chemical equilibria Based on these concentrations the proton condition can usually be simplified often leading directly to the value of the pH In the description of acid base titrations a general master equation is developed It provides a continuous and complete description of the entire titration curve which can then be used for computer based comparison with experimental data Graphical estimates of the steepness of titration curves are also developed from which the practicality of a given titration can be anticipated Activity effects are described in detail including their effect on titration curves The discussion emphasizes the distinction between equilibrium constants and electrometric pH measurements which are subject to activity corrections and balance equations and spectroscopic pH measurements which are not Finally an entire

chapter is devoted to what the pH meter measures and to the experimental and theoretical uncertainties involved **Acids, Bases and Salts MCQ (Multiple Choice Questions)** Arshad Iqbal, 2020 The Acids Bases and Salts Multiple Choice Questions MCQ Quiz with Answers PDF Acids Bases and Salts MCQ PDF Download Quiz Questions Practice Tests with Answer Key Class 10 Chemistry Questions Bank MCQs Notes includes revision guide for problem solving with solved MCQs Acids Bases and Salts MCQ with Answers PDF book covers basic concepts analytical and practical assessment tests Acids Bases and Salts MCQ PDF book helps to practice test questions from exam prep notes The Acids Bases and Salts MCQs with Answers PDF eBook includes revision guide with verbal quantitative and analytical past papers solved MCQs Acids Bases and Salts Multiple Choice Questions and Answers MCQs PDF Free download sample a book to practice quiz questions and answers on 10th grade chemistry topics What is acid base and salt acids and bases pH measurements self ionization of water pH scale Bronsted concept of acids and bases pH scale and salts tests for high school students and beginners Acids Bases and Salts Quiz Questions and Answers PDF free download eBook s sample covers exam s viva interview questions and competitive exam preparation with answer key The book Acids Bases and Salts MCQs PDF includes high school question papers to review practice tests for exams Acids Bases and Salts Multiple Choice Questions MCQ with Answers PDF digital edition eBook a study guide with textbook chapters tests for competitive exam Acids Bases and Salts Mock Tests eBook covers problem solving exam tests from high school chemistry textbooks

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