

Acids and Bases Worksheet

NAME _____

1. Name the following compounds as acids and circle which are weak acids:

- a)
- HNO_3
- b)
- H_2SO_4
- c)
- HF
- d)
- H_2CO_3
- e)
- $\text{HC}_2\text{H}_3\text{O}_2$
- f)
- H_3PO_4

2. Write the formulas for the following bases:

- a) calcium hydroxide b) aluminum hydroxide c) ammonium hydroxide d) lithium hydroxide

3. Complete the following table:

	Acids	Bases
Definition		
Electrolyte in aqueous solution?		
pH		
Taste		
Litmus turns		
Which reacts with metals to form H_2 gas?		

Conjugate acid-base pairs

4. Identify the acid, the base, the conjugate acid and the conjugate base in the reaction:



5. Identify the acid, the base, the conjugate acid and the conjugate base in the reaction:



6. An acid will form the conjugate _____ and a base will form the conjugate _____

Neutralization and Titration.

7. Classify the following acids as monoprotic, diprotic or triprotic:

- a)
- H_2SO_4
- b)
- H_3PO_4
- c)
- HCl
- d)
- $\text{HC}_2\text{H}_3\text{O}_2$
- or
- CH_3COOH
- (same thing)

8. What is the concentration of an unknown NaOH solution if took an average of 15.4mL of 0.100N HCl solution to neutralize 10.0mL of the NaOH solution?9. From the given data, what is the molarity and the normality of the HF solution if the KOH solution was 0.50M?

	Acid, HF	Base, KOH
Initial Volume, mL	5.5	14.9
Final Volume, mL	25.9	24.4

$\text{pH} = -\log[\text{H}_3\text{O}^+]$ To find pH if given $[\text{H}_3\text{O}^+]$, enter the $[\text{H}_3\text{O}^+]$, press Log & then the +/- button
To find $[\text{H}_3\text{O}^+]$ if given pH, enter negative pH, then 2nd function Log

10. A solution has a $[\text{H}_3\text{O}^+]$ of 1×10^{-7} M, what is its pH? _____ Is this an acid, base or is it neutral?

TURN THE PAPER OVER

11. What is the pH of a solution if the $[\text{H}^+] = 7.2 \times 10^{-6}\text{M}$? _____ Is this an acid or base?

Chapter 8 Solutions Acids And Bases Answer Key Test

Disha Experts



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Table of Contents Chapter 8 Solutions Acids And Bases Answer Key Test

1. Understanding the eBook Chapter 8 Solutions Acids And Bases Answer Key Test
 - The Rise of Digital Reading Chapter 8 Solutions Acids And Bases Answer Key Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 8 Solutions Acids And Bases Answer Key Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 8 Solutions Acids And Bases Answer Key Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 8 Solutions Acids And Bases Answer Key Test
 - Personalized Recommendations

- Chapter 8 Solutions Acids And Bases Answer Key Test User Reviews and Ratings
- Chapter 8 Solutions Acids And Bases Answer Key Test and Bestseller Lists
- 5. Accessing Chapter 8 Solutions Acids And Bases Answer Key Test Free and Paid eBooks
 - Chapter 8 Solutions Acids And Bases Answer Key Test Public Domain eBooks
 - Chapter 8 Solutions Acids And Bases Answer Key Test eBook Subscription Services
 - Chapter 8 Solutions Acids And Bases Answer Key Test Budget-Friendly Options
- 6. Navigating Chapter 8 Solutions Acids And Bases Answer Key Test eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 8 Solutions Acids And Bases Answer Key Test Compatibility with Devices
 - Chapter 8 Solutions Acids And Bases Answer Key Test Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 8 Solutions Acids And Bases Answer Key Test
 - Highlighting and Note-Taking Chapter 8 Solutions Acids And Bases Answer Key Test
 - Interactive Elements Chapter 8 Solutions Acids And Bases Answer Key Test
- 8. Staying Engaged with Chapter 8 Solutions Acids And Bases Answer Key Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 8 Solutions Acids And Bases Answer Key Test
- 9. Balancing eBooks and Physical Books Chapter 8 Solutions Acids And Bases Answer Key Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 8 Solutions Acids And Bases Answer Key Test
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 8 Solutions Acids And Bases Answer Key Test
 - Setting Reading Goals Chapter 8 Solutions Acids And Bases Answer Key Test
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 8 Solutions Acids And Bases Answer Key Test
 - Fact-Checking eBook Content of Chapter 8 Solutions Acids And Bases Answer Key Test

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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