Chapter 9 Stoichiometry Chapter Review Answers

- 1. Mole Statios
 - Mole ratios are conversion factors that relate the number of moles of one chemical to another chemical involved in a chemical reaction.
 - They are formed from the coefficients found in a balanced chemical equation.
- 2. Mole Ratios as conversion factors
 - $a. \quad \frac{2 \bmod Ca}{1 \bmod O_2}, \frac{1 \bmod O_2}{2 \bmod Ca}, \frac{2 \bmod Ca}{2 \bmod Ca}, \frac{2 \bmod Ca}{2 \bmod Ca}, \frac{1 \bmod O_2}{2 \bmod CaO}, \frac{1 \bmod O_2}{1 \bmod O_2}, \frac{2 \bmod CaO}{1 \bmod O_2}, \\ Remember the ratios can be reduced as well 11 \bmod O_2$
 - $\mathbf{b.} = \frac{1 \bmod Hg}{2 \bmod HF}, \frac{2 \bmod HF}{1 \bmod Hg}, \frac{1 \bmod HgF_2}{1 \bmod Hg}, \frac{1 \bmod Hg}{1 \bmod HgF_2}, \frac{1 \bmod Hg}{1 \bmod Hg}, \frac{1 \bmod Hg}{1 \bmod Hg}, \frac{2 \bmod HF}{1 \bmod Hg}, \frac{2 \bmod HF}{1 \bmod HgF_2}, \frac{2 \bmod HF}{2 \bmod HF}, \frac{1 \bmod HgF_2}{1 \bmod HgF_2}, \frac{1 \bmod HgF_2}, \frac{1 \bmod HgF_2}{1 \bmod HgF_2}, \frac{1 \bmod HgF_2}{1 \bmod HgF_2}, \frac{1 \bmod HgF_2}, \frac{1 \bmod HgF_2}{1 \bmod HgF_2}, \frac{1 \bmod HgF_2}{1 \bmod HgF_2}, \frac{1$
- $3. \quad \frac{165.79 \, g \, Na_{2}CO_{3}}{1 \, mol \, Na_{2}CO_{3}}, \\ \frac{1}{105.79 \, g \, Na_{2}CO_{3}}, \\ \frac{74.10 \, g \, Co(0H)_{2}}{1 \, mol \, Co(0H)_{2}}, \\ \frac{1}{74.10 \, g \, Co(0H)_{2}}, \\ \frac{1}{74.10 \, g \, Co(0H)_{2}}, \\ \frac{1}{100.01 \, g \, NaOM}, \\ \frac{1}{100.01 \, g \, CoCO_{3}}, \\ \frac{1}{100.01 \, g \, CO$
- 4. Molar Mass
 - a. Definition: the mass of one mole of a substance
 - Molar mass is used as a conversion factor in stoiciometric calculations to convert mass of a substance to moles of the substance or convert moles of a substance to mass of the substance.
- Hydrogen + Oxygen → Water
 - a. 5.0 mol H₂ b. 2.5 mol O₂
- 5. Ethane combustion reaction
 - a. 15.8 mol O₂
 b. 9.00 mol CO₂ & 13.5 mol H₂O
- 575 g Na & 886 g Cli
- 8. Inon production reaction
 - a. 75.1 mol CO
 b. 50.1 mol Fe & 75.1 mol CO₂
- 87,390 g CO or 8.739 × 10° g CO & 12,610 g H; or 1.261 × 10° g H;
- Lightning reaction of N₂ & O₂
 - a. 1,100 g NO₂ or 1.10 × 10³ g NO₃
- b. 720. g NO or 7.20 × 102 g NO

- 11. CO2 removal in spacecraft
 - a. 42.04 mol NaOH
- b. 21.02 mol Na₂CO₃ & 21.02 mol H₂O
- 12. Double replacement AgNO₃ & NaBr
 - a. 463 g NaBr
 b. 845 g AgBr
- 13. Soda-acid fire extinguisher
 - a. 3.058 mol NaHCO₃
- 3.058 mol CO₂, 1.529 mol Na₂SO₄, 3.058 mol H₂O
- 14. Sulfuric acid-sodium hydroxide neutralization reaction
 - a. H₂SO₀ + 2 NaOH → Na₂SO₀ + 2 H₂O
 - b. 37 g H₂SO₄
 - c. 53 g Na₁SO₁ & 14 g H₂O
- 15. Copper & silver nitrate reaction
 - a. 0.0104 mol Cu(NO₃)₂
 - b. 0.0104 mol Cu & 0.0209 mol AgNO₁
- 16. Aspirin production
 - a. 13.5 kg C₂H₂O₄
 - b. 7.66 kg C₄H₄O₃
 - 4.29 L HIC3H3O3
- 17. Ideal stoichiometry calculations do not account for factors that can affect the relative amounts of product chemicals formed in reactions. Ideal stoichiometric calculations are based on the amounts of reactants and products under ideal conditions. Real stoichiometric calculations account for actual conditions under which the reaction occurs.
- 18. The limiting reactant is the chemical in the reaction in least abundance for the reaction; it will be the chemical reactant that runs out first in the reaction. The limiting reactant fimits the amount of chemical product(s) that can be formed. The excess reactant(s) will not be used up completely in the reaction; they will be leftover.

Chapter 9 Stoichiometry Worksheet Answers

Michael Seilmaier

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Holt Chemistry Ralph Thomas Myers, 2004 Objective Workbook for Simplified ICSE Chemistry, Books in Print Supplement ,2002 Merrill Chemistry Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 **Spreadsheet Chemistry** O. Jerry Parker, Gary L. Breneman, 1991 Chemical Engineering Design Gavin Towler, Ray Sinnott, 2012-01-25 Chemical Engineering Design Second Edition deals with the application of chemical engineering principles to the design of chemical processes and equipment Revised throughout this edition has been specifically developed for the U S market It provides the latest US codes and standards including API ASME and ISA design codes and ANSI standards It contains new discussions of conceptual plant design flowsheet development and revamp design extended coverage of capital cost estimation process costing and economics and new chapters on equipment selection reactor design and solids handling processes A rigorous pedagogy assists learning with detailed worked examples end of chapter exercises plus supporting data and Excel spreadsheet calculations plus over 150 Patent References for downloading from the companion website Extensive instructor resources including 1170 lecture slides and a fully worked solutions manual are available to adopting instructors This text is designed for chemical and biochemical engineering students senior undergraduate year plus appropriate for capstone design courses where taken plus graduates and lecturers tutors and professionals in industry chemical process biochemical pharmaceutical petrochemical sectors New to this edition Revised organization into Part I Process Design and Part II Plant Design The broad themes of Part I are flowsheet development economic analysis safety and environmental impact and optimization Part II contains chapters on equipment design and selection that can be used as supplements to a lecture course or as essential references for students or practicing engineers working on design projects New discussion of conceptual plant design flowsheet development and revamp design Significantly increased coverage of capital cost estimation process costing and economics New chapters on equipment selection reactor design and solids handling processes New sections on fermentation adsorption membrane separations ion exchange and chromatography Increased coverage of batch processing food pharmaceutical and biological processes All equipment chapters in Part II revised and updated with current information Updated throughout for latest US codes and standards including API ASME and ISA design codes and ANSI standards Additional worked examples and homework problems The most complete and up to date coverage of equipment selection 108 realistic commercial design projects from diverse industries A rigorous pedagogy assists learning with detailed worked examples end of chapter exercises plus supporting data and Excel spreadsheet calculations plus over 150 Patent References for downloading from the companion website Extensive instructor resources 1170 lecture slides plus fully worked solutions manual available to adopting instructors

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