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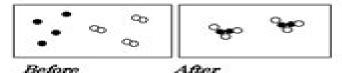
Chemistry - Balancing Equations

Balance the following equations by inserting the proper coefficients. For selected reactions, draw *Before* and *After* particle diagrams to show the particles involved in the reaction. Be sure to provide a key.

10.3

#18

- 1. $\underline{1}$ C + $\underline{1}$ H₂O \rightarrow $\underline{1}$ CO + $\underline{1}$ H₂
- 2. $2 \text{ MgO} \rightarrow 2 \text{ Mg} + 1 \text{ O}_2$
- 3. _4_Al + _3_O₂ → _2_Al₂O₃



- 4. $1 \text{ Zn} + 1 \text{ H}_2\text{SO}_4 \rightarrow 1 \text{ ZnSO}_4 + 1 \text{ H}_2$
- 5. $1 \text{ Cl}_2 + 2 \text{ KI} \rightarrow 2 \text{ KCl} + 1 \text{ I}_2$
- 6. 2 CuCl \rightarrow 2 Cu + 1 Cl₂
- 2 Na + 1 Cl₂ → 2 NaCl
- 8. $2\text{Al} + 6\text{HCl} \rightarrow 2\text{AlCl}_3 + 3\text{H}_2$



- 9. $2 \text{ Fe}_2\text{O}_3 \rightarrow 4 \text{ Fe} + 3 \text{ O}_2$
- 10. $\underline{4} P + \underline{5} O_2 \rightarrow \underline{2} P_2 O_5$
- 11. $\underline{1}$ Mg + $\underline{2}$ HCl \rightarrow $\underline{1}$ MgCl₂ + $\underline{1}$ H₂
- 12. $3 \text{ H}_2 + 1 \text{ N}_2 \rightarrow 2 \text{ NH}_3$
- 13. $\underline{1}$ BaCl₂ + $\underline{1}$ H₂SO₄ \rightarrow $\underline{1}$ BaSO₄ + $\underline{2}$ HCl
- 14. <u>1 CH4 + 2 O2 → 1 CO2 + 2 H2O</u>



#14 Before After

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Kenneth W. Raymond

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Survival Guide to General Chemistry Patrick E. McMahon, Rosemary McMahon, Bohdan Khomtchouk, 2019-02-13 This work evolved over thirty combined years of teaching general chemistry to a variety of student demographics The focus is not to recap or review the theoretical concepts well described in the available texts Instead the topics and descriptions in this book make available specific detailed step by step methods and procedures for solving the major types of problems in general chemistry Explanations instructional process sequences solved examples and completely solved practice problems are greatly expanded containing significantly more detail than can usually be devoted to in a comprehensive text Many chapters also provide alternative viewpoints as an aid to understanding Key Features The authors have included every major topic in the first semester of general chemistry and most major topics from the second semester Each is written in a specific and detailed step by step process for problem solving whether mathematical or conceptual Each topic has greatly expanded examples and solved practice problems containing significantly more detail than found in comprehensive texts Includes a chapter designed to eliminate confusion concerning acid base reactions which often persists through working with acid base equilibrium Many chapters provide alternative viewpoints as an aid to understanding This book addresses a very real need for a large number of incoming freshman in STEM fields General Organic and Biological Chemistry Kenneth W. Raymond, 2009-12-14 This general organic and biochemistry text has been written for students preparing for careers in health related fields such as nursing dental hygiene nutrition medical technology and occupational therapy It is also suited for students majoring in other fields where it is important to have an understanding of the basics of chemistry Students need have no previous background in chemistry but should possess basic math skills The text features numerous helpful problems and learning features The Practice of Chemistry Donald J. Wink, Sharon Fetzer-Gislason, Sheila McNicholas, 2003-03 Students can t do chemistry if they can t do the math The Practice of Chemistry First Edition is the only preparatory chemistry text to offer students targeted consistent mathematical support to make sure they understand how to use math especially algebra in chemical problem solving The book s unique focus on actual chemical practice extensive study tools and integrated media makes The Practice of Chemistry the most effective way to prepare students for the standard general chemistry course and bright futures as science majors This special PowerPoint tour of the text was created by Don Wink http www bfwpub com pdfs wink POCPowerPoint Final ppt 832KB Chemistry Neil D. Jespersen, Alison Hyslop, 2021-11-02 Chemistry The Molecular Nature of Matter 8th Edition continues to focus on the intimate relationship that exists between structure at the atomic molecular level and the observable macroscopic properties of matter Key revisions in this edition focus on three areas The deliberate inclusion of more updated real world examples that relate common real world student experiences to the science of chemistry Simultaneously examples and questions have been updated to align them with career concepts relevant to the environmental engineering biological pharmaceutical and medical sciences Providing students with transferable skills

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