



ELECTROCHEMISTRY

1. ELECTROCHEMISTRY

Electrochemistry is the study of production of electricity from the energy released during a spontaneous chemical reaction and the use of electrical energy to bring about non-spontaneous chemical transformations.

2. ELECTROCHEMICAL CELLS

A spontaneous chemical process is the one which can take place on its own and in such a process the Gibbs energy of the system decreases. It is this energy that gets converted to electrical energy. The reverse process is also possible in which we can make non-spontaneous processes occur by supplying external energy in the form of electrical energy. These inter conversions are carried out in equipments called Electrochemical Cells.

3. TYPES

Electrochemical Cells are of two types:

3.1 Galvanic Cells

Converts chemical energy into electrical energy

3.2 Electrolytic Cells

Converts electrical energy into chemical energy.

4. GALVANIC CELL

Cell energy is extracted from a spontaneous chemical process or reaction and it is converted to electric current. For example, Daniell Cell is a Galvanic Cell in which Zinc and Copper are used for the redox reaction to take place.



Zn is the reducing agent and Cu^{2+} is the oxidising agent. The half cells are also known as **Electrodes**. The oxidation half is known as **Anode** and the reduction half is called **Cathode**. Electrons flow from anode to cathode in

the external circuit. Anode is assigned **negative polarity** and cathode is assigned **positive polarity**. In Daniell Cell, Zn acts as the anode and Cu acts as the cathode.

5. ELECTROLYTIC CELL

These electrodes are dipped in an electrolytic solution containing cations and anions. On supplying current the ions move towards electrodes of opposite polarity and simultaneous reduction and oxidation takes place.

5.1 Preferential Discharge of ions

Where there are more than one cation or anion the process of discharge becomes competitive in nature. Discharge of any ion requires energy and in case of several ions being present the discharge of that ion will take place first which requires the energy.

6. ELECTRODE POTENTIAL

It may be defined as the tendency of an element, when it is placed in contact with its own ions to either lose or gain electrons and in turn become positively or negatively charged.

The electrode potential will be named as oxidation or reduction potential depending upon whether oxidation or reduction has taken place.



6.1 Characteristics

- (a) Both oxidation and reduction potentials are equal in magnitude but opposite in sign.
- (b) It is not a thermodynamic property, so values of E are not additive.

7. STANDARD ELECTRODE POTENTIAL (E°)

It may be defined as the electrode potential of an electrode determined relative to standard hydrogen electrode under standard conditions. The standard conditions taken are :



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Table of Contents Chapter Chemistry Study Guide Answers

1. Understanding the eBook Chapter Chemistry Study Guide Answers
 - The Rise of Digital Reading Chapter Chemistry Study Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter Chemistry Study Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter Chemistry Study Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter Chemistry Study Guide Answers
 - Personalized Recommendations
 - Chapter Chemistry Study Guide Answers User Reviews and Ratings

- Chapter Chemistry Study Guide Answers and Bestseller Lists
- 5. Accessing Chapter Chemistry Study Guide Answers Free and Paid eBooks
 - Chapter Chemistry Study Guide Answers Public Domain eBooks
 - Chapter Chemistry Study Guide Answers eBook Subscription Services
 - Chapter Chemistry Study Guide Answers Budget-Friendly Options
- 6. Navigating Chapter Chemistry Study Guide Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter Chemistry Study Guide Answers Compatibility with Devices
 - Chapter Chemistry Study Guide Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter Chemistry Study Guide Answers
 - Highlighting and Note-Taking Chapter Chemistry Study Guide Answers
 - Interactive Elements Chapter Chemistry Study Guide Answers
- 8. Staying Engaged with Chapter Chemistry Study Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter Chemistry Study Guide Answers
- 9. Balancing eBooks and Physical Books Chapter Chemistry Study Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter Chemistry Study Guide Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter Chemistry Study Guide Answers
 - Setting Reading Goals Chapter Chemistry Study Guide Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter Chemistry Study Guide Answers
 - Fact-Checking eBook Content of Chapter Chemistry Study Guide Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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