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E-commerce 2023–2024

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EIGHTEENTH EDITION

Kenneth C. Laudon
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web jan 29 2020 the ionic form of this reaction is 2 na aq 2 cl aq cu 2 aq so₄²⁻ aq 2 na aq so₄²⁻ aq cucl₂ s the sodium ions and sulfate ion are the spectator ions in this reaction they appear unchanged in both the product and reactant side of the equation

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web the six no₃⁻ aq ions and the six na aq ions that appear on both sides of the equation are spectator ions that can be canceled to give the net ionic equation 3ba 2 aq 2po₄³⁻ aq → ba₃(po₄)₂ s

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web these spectator ions whose presence is required to maintain charge neutrality are neither chemically nor physically changed by the process and so they may be eliminated from the equation to yield a more succinct representation called a

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web the equation below is the final net ionic equation after elimination of the spectator ions which yields water $\text{H}^+ \text{aq} + \text{OH}^- \text{aq} \rightarrow \text{H}_2\text{O} \text{l}$ the spectator ions in $\text{KOH} \text{aq}$ and $\text{HNO}_3 \text{aq}$ are $\text{K}^+ \text{aq}$ and $\text{NO}_3^- \text{aq}$ how to identify and list the

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web the $\text{Ag}^+ \text{aq}$ and $\text{Cl}^- \text{aq}$ ions become $\text{AgCl} \text{s}$ but the $\text{Na}^+ \text{aq}$ ions and the $\text{NO}_3^- \text{aq}$ ions stay as $\text{Na}^+ \text{aq}$ ions and $\text{NO}_3^- \text{aq}$ ions these two ions are examples of spectator ions ions that do nothing in the overall course of a chemical reaction

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