Concentration and Mol:	arity PhET Labs
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Hame

is added and when

Procedure: https://phet.colorado.edu/hims/html/concentration/latest/concentration_en.html

Part 1: Dissolution and Saturation

Take some time to play and familiarize yourself with the simulation. Click on everything. Move all

....



Concentration

evaporation occurs.

How does the concentration change as solid solute is added?

2. How does the concentration change as additional water is added? concentration decreases

0.179

Concentration

3. How does the concentration change as evaporation occurs? concentration increased

4. How do you know when a solution is soluroted? concentration does not change (it also shows on screen)

6. When a solution is saturated, and additional solid solute is added, what happens? Solid does not dissolve and settles at bottom

7. Why do you think this is? solvent cannot hold any more solute.

the sliders. Notice what happens to the concentration as solid solute

8. How does adding this additional solute change the concentration of this saturated solution? decreases concentration

9. How does evaporation change the concentration of a saturated solution? 80 change solution is already saturated

Part 2: Concentrated Solutions (same website as part 1 kmt click on the Solution option)

It Solution -

10. Adding a concentrated solution ... describe a way to determine the concentration of the solution in the spigot. Write your plan heres _empty the container and add only one solution from the dropper and then use the

sensor to measure concentration

Desired Desired

Count distings Calculations and Pulsanium districts and

Course where

Parameter persons per

11. Using your plan...how might you get that concentrated solution to become saturated? You can evaporate

12. Does your plan work for all the other solutions too? Yes Why? / Why Not? evaporation reduces the

amount of water (or solute) so concentration increases

Part 3: Molarity https://phet.colorada.edu/sims/html/molarity/latest/molarity_en.html

Malaray (M) = amount of solver (mol) related of solvers (L)

Molarity is moles per Liter, that is, how many moles of solute (entire salt) is dissolved per Liter of solution.

First, determine the saturation concentration of each of the solutions, that is, how concentrated can you get each solution before the solution is saturated. If you can't determine the concentration using the simulation "Molarity", try using the simulation "Concentration" (You will use this information again in Port S. If your instructor requires it)



Molarity

Cobalt (II) nitrate	5.640 M	Potassium chromate	3.350 M
Cobalt chloride	4.330 M	Nickel (II) chloride	5.210 M
Potassium dichromate	0.510 M	Copper sulfate	1.380 M
Gold (III) chloride	2.30 M	Potassium permanganate	0.480 M

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Gary D. Christian

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