

Practice makes perfect with Faraday's Law  
Using conversion factors in Electrochemistry

1. Determine the mass of copper that is plated out when a current of 10.0 amps is passed for 30.0 minutes through a solution containing  $\text{Cu}^{2+}$

$$\text{Cu}^{+2} + 2\text{e}^- \rightarrow \text{Cu}(s)$$

$$30 \text{ min} \times \frac{60 \text{ s}}{1 \text{ min}} = 1800 \text{ s} \quad / \quad \frac{1 \text{ mole e}^-}{96485 \text{ C}} \times \frac{1 \text{ mole Cu}}{2 \text{ mole e}^-} \times \frac{63.55 \text{ g}}{1 \text{ mole Cu}}$$

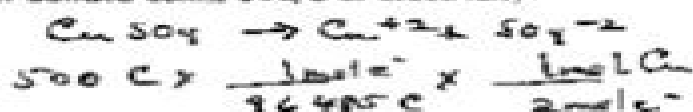
5.83 g Answer

2. How long will it take to plate out 1.0 kg Al from aqueous Al using a current of 100.0 A? Calculate in minutes

$$1000 \text{ g Al} \times \frac{1 \text{ mole Al}}{26.98 \text{ g}} \times \frac{3 \text{ mole e}^-}{1 \text{ mole Al}} \times \frac{96485 \text{ C}}{1 \text{ mole e}^-} \times \frac{1 \text{ sec}}{1000 \text{ C}} \times \frac{1 \text{ min}}{60 \text{ s}}$$

1788 min Answer 1800 min

3. Calculate the moles of copper metal that can be produced by the electrolysis of molten copper sulfate using 500 C of electricity.



0.00259 or 2.59 x 10<sup>-3</sup> Answer moles

4. What mass of copper could be deposited from a copper(II) sulfate solution using a current of 0.50 A over 10 seconds?

$$\text{CuSO}_4 \rightarrow \text{Cu}^{+2} + \text{SO}_4^{-2}$$

$$.50 \frac{\text{C}}{\text{s}} \times 10 \text{ s} \times \frac{1 \text{ mole e}^-}{96485 \text{ C}} \times \frac{1 \text{ mole Cu}}{2 \text{ mole e}^-} \times \frac{63.55 \text{ g}}{1 \text{ mole Cu}}$$

1.6 x 10<sup>-3</sup> g Answer

5. Calculate the time required to deposit 56 g of silver from a silver nitrate solution using a current of 4.5 A.



11631 s = 11000 s Answer  
3.1 hrs

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**Guillaume Favre**



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## **Table of Contents Concept Development 37 1 Faraday S Law Answer Key**

1. Understanding the eBook Concept Development 37 1 Faraday S Law Answer Key
  - The Rise of Digital Reading Concept Development 37 1 Faraday S Law Answer Key
  - Advantages of eBooks Over Traditional Books
2. Identifying Concept Development 37 1 Faraday S Law Answer Key
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Concept Development 37 1 Faraday S Law Answer Key
  - User-Friendly Interface
4. Exploring eBook Recommendations from Concept Development 37 1 Faraday S Law Answer Key
  - Personalized Recommendations
  - Concept Development 37 1 Faraday S Law Answer Key User Reviews and Ratings
  - Concept Development 37 1 Faraday S Law Answer Key and Bestseller Lists
5. Accessing Concept Development 37 1 Faraday S Law Answer Key Free and Paid eBooks
  - Concept Development 37 1 Faraday S Law Answer Key Public Domain eBooks

- Concept Development 37 1 Faraday S Law Answer Key eBook Subscription Services
- Concept Development 37 1 Faraday S Law Answer Key Budget-Friendly Options
- 6. Navigating Concept Development 37 1 Faraday S Law Answer Key eBook Formats
  - ePub, PDF, MOBI, and More
  - Concept Development 37 1 Faraday S Law Answer Key Compatibility with Devices
  - Concept Development 37 1 Faraday S Law Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Concept Development 37 1 Faraday S Law Answer Key
  - Highlighting and Note-Taking Concept Development 37 1 Faraday S Law Answer Key
  - Interactive Elements Concept Development 37 1 Faraday S Law Answer Key
- 8. Staying Engaged with Concept Development 37 1 Faraday S Law Answer Key
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Concept Development 37 1 Faraday S Law Answer Key
- 9. Balancing eBooks and Physical Books Concept Development 37 1 Faraday S Law Answer Key
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Concept Development 37 1 Faraday S Law Answer Key
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Concept Development 37 1 Faraday S Law Answer Key
  - Setting Reading Goals Concept Development 37 1 Faraday S Law Answer Key
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Concept Development 37 1 Faraday S Law Answer Key
  - Fact-Checking eBook Content of Concept Development 37 1 Faraday S Law Answer Key
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks

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#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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