# Thermodynamics

- 1. Which of the following statements is correct for the spontaneous absorption of a gas?
  - (a) AS is negative and therefore, AH should be highly positive
  - (b) AS is negative and therefore, AH should be highly negative
  - (c) AS is positive and therefore, AH should be
  - (d) AS is positive and therefore, AH should also be highly positive
- For the reaction, X<sub>2</sub>O<sub>2</sub>(I) -> 2XO<sub>2</sub>(g).

 $\Delta U = 2.1 \text{ kcal}$ ,  $\Delta S = 20 \text{ cal K}^{-1} \text{ at } 300 \text{ K}$ . Hence, AG is **[2014]** 

(a) 2.7 kcal (c) 9.3 keal

(b) - 2.7 kcal (d) = 9.3 kcall

3. For a given exothermic reaction,  $K_p$  and  $K'_p$  are the equilibrium constants at temperatures T, and T<sub>2</sub>, respectively. Assuming that heat of reaction is constant in temperature range between  $T_1$  and  $T_2$ , it is readily observed that

(a)  $K_{\alpha} > K_{\alpha}^*$ 

- (b)  $K_0 < K_0$
- (c)  $K_{\phi} = K_{\phi}^*$
- (d)  $K_{p} = \frac{1}{K_{p}^{*}}$
- 4. A reaction having equal energies of activation for forward and reverse reactions
  - (a)  $\Delta S = 0$

(b)  $\Delta G = 0$ 

(c)  $\Delta H = 0$ 

(d)  $\Delta H = \Delta G = \Delta S = 0$ 

- 5. In which of the following reactions standard reaction entropy changes  $(\Delta S^{\circ})$  is positive and standard Gibbs energy change (AG°) decreases sharply with increasing temperature? [2012]
  - (a) C (graphite) +  $\frac{1}{2}O_2(g) \longrightarrow CO(g)$
  - (b)  $CO(g) + \frac{1}{2}O_2(g) \longrightarrow CO_2(g)$
  - (c)  $Mg(s) + \frac{1}{2}O_2(g) \longrightarrow MgO(s)$
  - (d)  $\frac{1}{2}$  C (graphite) +  $\frac{1}{2}$  O<sub>2</sub>(g)  $\longrightarrow \frac{1}{2}$  CO<sub>2</sub>(g)
- 6. The enthalpy of fusion of water is 1.435 kcal/mol. The molar entropy change for the melting of ice at 0°C is
  - (a) 10:52 cal/mol K.

(b) 21.04 cal/mol K

- (c) 5.260 cal/mol K
- (d) 0.526 cal/mol K
- Standard enthalpy of vaporisation \(\Delta\_{\text{cop}} H^\*\) for water at 100°C is 40.66 k] mol-1. The internal energy of vaporisation of water at 100°C (in kJ mol-1) is (assume water vapour to behave like an ideal gas). [2012]
  - (a) + 37.56
  - (c) + 43.76

- (b) 43.76(d) + 40.66
- If the enthalpy change for the transition of liquid water to steam is 30 kl mol<sup>-1</sup> at 27°C. the entropy change for the process would be
  - (a) 1.0 J moi\*\*K-1

[2011]

- (b) 0.1 J mol<sup>-1</sup>k<sup>-1</sup>
- (c) 100 J mei<sup>-1</sup>k<sup>-1</sup>
- (d) 10 J mot 1k-1

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**R Barnett** 

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#### **Table of Contents Conceptual Physics Thermodynamics Answers**

- 1. Understanding the eBook Conceptual Physics Thermodynamics Answers
  - The Rise of Digital Reading Conceptual Physics Thermodynamics Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Conceptual Physics Thermodynamics Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Conceptual Physics Thermodynamics Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Conceptual Physics Thermodynamics Answers

- Personalized Recommendations
- Conceptual Physics Thermodynamics Answers User Reviews and Ratings
- Conceptual Physics Thermodynamics Answers and Bestseller Lists
- 5. Accessing Conceptual Physics Thermodynamics Answers Free and Paid eBooks
  - Conceptual Physics Thermodynamics Answers Public Domain eBooks
  - Conceptual Physics Thermodynamics Answers eBook Subscription Services
  - Conceptual Physics Thermodynamics Answers Budget-Friendly Options
- 6. Navigating Conceptual Physics Thermodynamics Answers eBook Formats
  - o ePub, PDF, MOBI, and More
  - Conceptual Physics Thermodynamics Answers Compatibility with Devices
  - Conceptual Physics Thermodynamics Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Conceptual Physics Thermodynamics Answers
  - Highlighting and Note-Taking Conceptual Physics Thermodynamics Answers
  - o Interactive Elements Conceptual Physics Thermodynamics Answers
- 8. Staying Engaged with Conceptual Physics Thermodynamics Answers
  - o Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Conceptual Physics Thermodynamics Answers
- 9. Balancing eBooks and Physical Books Conceptual Physics Thermodynamics Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Conceptual Physics Thermodynamics Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Conceptual Physics Thermodynamics Answers
  - Setting Reading Goals Conceptual Physics Thermodynamics Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Conceptual Physics Thermodynamics Answers

- Fact-Checking eBook Content of Conceptual Physics Thermodynamics Answers
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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