

3. a) The mean kinetic energy of a nitrogen molecule is the same as the mean kinetic energy of an oxygen molecule since E_k is directly proportional to temperature $E_k = \frac{3}{2} kT$

b) molar mass $N_2 = 0.028 \text{ kg mol}^{-1}$ molar mass $O_2 = 0.032 \text{ kg mol}^{-1}$
 $N_A = 6.02 \times 10^{23} \text{ mol}^{-1}$

$$\frac{1}{2} m_N c_{\text{rms},N}^2 = \frac{1}{2} m_O c_{\text{rms},O}^2$$

$$\frac{0.028}{6.02 \times 10^{23}} c_{\text{rms},N}^2 = \frac{0.032}{6.02 \times 10^{23}} c_{\text{rms},O}^2$$

$$0.028 c_{\text{rms},N}^2 = 0.032 c_{\text{rms},O}^2$$

$$c_{\text{rms},N}^2 = \frac{0.032}{0.028} c_{\text{rms},O}^2$$

$$c_{\text{rms},N}^2 = 1.14 c_{\text{rms},O}^2$$

$$c_{\text{rms},N} = \sqrt{1.14} c_{\text{rms},O}$$

$$c_{\text{rms},N} = 1.07 c_{\text{rms},O}$$

4. molar mass = $0.028 \text{ kg mol}^{-1}$ $V = 0.037 \text{ m}^3$ $T = 300 \text{ K}$
 $p = 100 \text{ kPa} = 100 \times 10^3 \text{ Pa}$

a) $pV = nRT \Rightarrow \frac{pV}{RT} = n = \frac{(100 \times 10^3)(0.037)}{8.31(300)}$
 $= \frac{3700}{2493} = 1.484 = \underline{1.48 \text{ moles to 3sf}}$

b) mass = number of moles $\times$ molar mass
 $= 1.484 \times 0.028 = 0.041512 \text{ kg}$
 $= \underline{4.2 \times 10^{-2} \text{ kg to 2sf}}$

(If 1.48 moles used then mass = $0.04144 \text{ kg} = 4.1 \times 10^{-2} \text{ kg to 2sf}$)

c) mass of one molecule $m = \frac{\text{molar mass}}{N_A} = \frac{0.028}{6.02 \times 10^{23}} = 4.65 \times 10^{-26} \text{ kg}$

$$\frac{1}{2} m c_{\text{rms}}^2 = \frac{3}{2} kT$$

$$4.65 \times 10^{-26} c_{\text{rms}}^2 = 3 \left(1.38 \times 10^{-23} \right) 300 = 1.242 \times 10^{-20}$$

$$c_{\text{rms}} = \frac{1.242 \times 10^{-20}}{4.65 \times 10^{-26}} = 267 \times 10^3 \text{ m s}^{-1}$$

$$c_{\text{rms}} = \sqrt{267 \times 10^3}$$

$$= 516 \text{ m s}^{-1}$$

$$= \underline{5.2 \times 10^2 \text{ m s}^{-1} \text{ to 2sf}}$$

n.b could use
 $pV = \frac{1}{2} N m c_{\text{rms}}^2$
 Since $N m = \text{mass}$
 and would also be
 equal to 1.48(0.028)

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