

Part C: Determination of the molar mass of naphthaleneName of solute: CyclohexaneMass of solute: 50 g

Freezing points

Mass of solvent: 7.71 gTrial 1: -3.3 °CTrial 2: -3.5 °CTrial 3: -3.5 °CAverage: -3.43 °C

1. Calculate the freezing point depression caused by the addition of naphthalene.

$$-10.36$$

2. Calculate the molality of the cyclohexane/naphthalene solution. Use the freezing point depression constant calculated from Part B.

$$\Delta T_f = K_f \times M$$

$$\frac{-10.36}{-18.35} = \frac{+18.35 \times M}{-18.35}$$

$$M = .565$$

$$K_f = -18.35 \quad \Delta T_f = -10.36$$

3. How many moles of the solute are present? What is the experimental molar mass of the solute? Use your equations for molality and freezing point depression to solve for this.

$$\Delta T_f = K_f \frac{\text{moles solute}}{\text{kg solvent}}$$

$$\frac{.5 g}{.00771 kg} = 116.21$$

$$-10.36 \div -18.35 = .0043 \text{ mol}$$

4. The molar mass for naphthalene is 128.17 g/mole. Calculate the percent error between your experimental molar mass and the actual molar mass.

$$\left| \frac{T - E}{T} \right| \times 100$$

$$\left(\frac{128.17 - 116.21}{128.17} \right) \times 100 = 9.28\% \text{ error}$$

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Minjie Lin



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