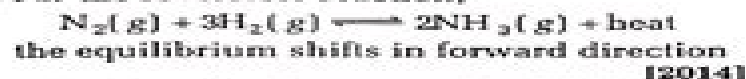


7

Chemical Equilibrium

1. For the reversible reaction,



- (a) by increasing the concentration of $\text{NH}_3(\text{g})$
 (b) by decreasing the pressure
 (c) by decreasing the concentrations of $\text{N}_2(\text{g})$ and $\text{H}_2(\text{g})$
 (d) by increasing pressure and decreasing temperature

2. Using the Gibbs energy change, $\Delta G^\circ = +63.3 \text{ kJ}$ for the following reaction,



the K_{sp} of $\text{Ag}_2\text{CO}_3(\text{s})$ in water at 25°C is
 ($R = 8.314 \text{ J K}^{-1} \text{ mol}^{-1}$) [2014]

- (a) 3.2×10^{-26}
 (b) 8.0×10^{-12}
 (c) 2.9×10^{-3}
 (d) 7.9×10^{-2}

3. KMnO_4 can be prepared from K_2MnO_4 as per reaction,



The reaction can go to completion by removing OH^- ions by adding [2013]

- (a) HCl
 (b) KOH
 (c) CO_2
 (d) SO_2

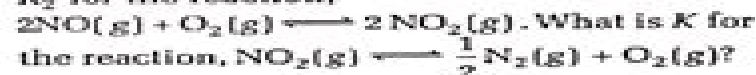
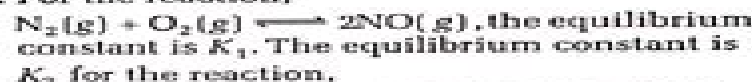
4. The value of ΔH for the reaction,



is less than zero. Formation of $\text{XY}_4(\text{g})$ will be favoured at [2011]

- (a) low pressure and low temperature
 (b) high temperature and low pressure
 (c) high pressure and low temperature
 (d) high temperature and high pressure

5. For the reaction,



- (a) $1/(4K_1K_2)$ (b) $\{1/K_1K_2\}^{1/2}$ [2011]
 (c) $1/(K_1K_2)$ (d) $1/4K_1K_2$

6. In which of the following equilibrium K_c and K_p are not equal? [2010]

- (a) $2\text{NO}(\text{g}) \rightleftharpoons \text{N}_2(\text{g}) + \text{O}_2(\text{g})$
 (b) $\text{SO}_2(\text{g}) + \text{NO}_2(\text{g}) \rightleftharpoons \text{SO}_3(\text{g}) + \text{NO}(\text{g})$
 (c) $\text{H}_2(\text{g}) + \text{I}_2(\text{g}) \rightleftharpoons 2\text{HI}(\text{g})$
 (d) $2\text{C}(\text{s}) + \text{O}_2(\text{g}) \rightleftharpoons 2\text{CO}_2(\text{g})$

7. The dissociation constants for acetic acid and HCN at 25°C are 1.5×10^{-5} and 4.5×10^{-10} , respectively. The equilibrium constant for the equilibrium,



- (a) 3.0×10^5 (b) 3.0×10^{-5}
 (c) 3.0×10^{-4} (d) 3.0×10^4

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Frederick William Miller



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