

EXPERIMENT 2: Analysis of Ions in Commercial Preparations

DATE: September 8, 2023

A. Results

1) Alum

a) Physical Tests:

- i) Color: White
- ii) Physical State: Fine crystalline powder

b) Chemical Tests:

a.) 1. A white precipitate is formed upon adding 1mL of 1N NaOH.

2. The white precipitate dissolves upon adding excess 1N NaOH but the color of the sample changed to a clear yellowish substance. (Aluminum Ion is present and $\text{Al}(\text{OH})_4$ is formed)

b.) 1. In addition of sodium bitartrate to the alum solution there is no precipitate formed.

2. There's still no precipitate formed after adding excess of Sodium Bitartrate but the color of the sample change to a very light greenish substance.

c.) 1. A clear white precipitate is formed upon adding 6N Ammonium Hydroxide.

2. An excessive amount of ammonium hydroxide produced a precipitate that was still clear and white but did not dissolve.

d.) 1. The alum solution formed a milky white cloud precipitate. (Sulfate ion is present)

2.

Ion(s) Present: Aluminum Ion (Al^{+3}), sulfate ion (SO_4^{2-})

2) Baking Soda

a) Physical Tests:

- i) Color: White (powder)
- ii) Physical State: odorless fine powdered solid matter

b) Chemical Tests:

A. N/A

B. Small bubbles gradually rising up with the addition of 1 mL of 1N sulfuric acid; many bubbles quickly rising up along with a white precipitate, in the addition of 1 mL of calcium hydroxide T.S.

C. N/A

c. Ion(s) present: Sodium ion (Na^+) and Bicarbonate ion (HCO_3^-)

3) Laundry Bleaching Agent

a) Physical Tests:

- i) Color: Yellow transparent
- ii) Physical State: Liquid

b) Chemical Test:

- i) Red litmus paper turns to pale blue which signifies the substance as base (alkaline).
- ii) It effervesces, and the color of the combined substance is transparent yellow.

4) Calamine Lotion

a) Physical Tests:

- i) Color: Light cream orange

Chem Lab Manual Answers

Ralph A. Burns



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Chemistry Lab Manual Answer Key, 2nd Edition Bright Thinker, 2024-07 Organic Chemistry I Laboratory Matt Davies, 2015-11-11 Laboratory Manual for General Chemistry Alvin Strickler, 1923 **Exploring General Chemistry in the Laboratory** Colleen F. Craig, Kim N. Gunnerson, 2017-02-01 This laboratory manual is intended for a two semester general chemistry course The procedures are written with the goal of simplifying a complicated and often challenging subject for students by applying concepts to everyday life This lab manual covers topics such as composition of compounds reactivity stoichiometry limiting reactants gas laws calorimetry periodic trends molecular structure spectroscopy kinetics equilibria thermodynamics electrochemistry intermolecular forces solutions and coordination complexes By the end of this course you should have a solid understanding of the basic concepts of chemistry which will give you confidence as you embark on your career in science **Basic Concepts of Chemistry** Leo J. Malone, Theodore Dolter, 2008-12-03 Engineers who need to have a better understanding of chemistry will benefit from this accessible book It places a stronger emphasis on outcomes assessment which is the driving force for many of the new features Each section focuses on the development and assessment of one or two specific objectives Within each section a specific objective is included an anticipatory set to orient the reader content discussion from established authors and guided practice problems for relevant objectives These features are followed by a set of independent practice problems The expanded Making it Real feature showcases topics of current interest relating to the subject at hand such as chemical forensics and more medical related topics Numerous worked examples in the text now include Analysis and Synthesis sections which allow engineers to explore concepts in greater depth and discuss outside relevance **Integrated Physics and Chemistry Lab Manual Answer Key, 3rd Edition** Bright Thinker, 2024-08 Key Individual Lab Manual Answer Key for Integrated Physics and Chemistry IPC 3rd edition **Laboratory Manual for Principles of General Chemistry** J. A. Beran, Mark Lassiter, 2022-08-16 Laboratory Manual for Principles of General Chemistry 11th Edition covers two semesters of a general chemistry laboratory program The material focuses on the lab experiences that reinforce the concepts that not all experimental conclusions are the same and depend on identifying an appropriate experimental procedure selecting the proper apparatus employing the proper techniques systematically analyzing and interpreting the data and minimizing inherent variables As a result of good data a scientific and analytical conclusion is made which may or may not be right but is certainly consistent with the data Experiments write textbooks textbooks don't write experiments A student's scientific literacy grows when experiences and observations associated with the scientific method are encountered Further experimentation provides additional cause effect observations leading to an even better understanding of the experiment The 11th edition's experiments are informative and challenging while offering a solid foundation for technique safety and experimental procedure The reporting and analysis of the data and the pre and post lab questions focus on the intuitiveness of the experiment The experiments may accompany any general chemistry textbook

and are compiled at the beginning of each curricular unit An Additional Notes column is included in each experiment s Report Sheet to provide a space for recording observations and data during the experiment Continued emphasis on handling data is supported by the Data Analysis section

LAB MANUAL FOR CHEMISTRY: ATOMS FIRST Gregg Dieckmann, John W. Sibert, 2017-10-31 This laboratory manual presents a curriculum that is organized around an atoms first approach to general chemistry Our motivation for writing this manual is to 1 tap into the natural curiosity present in all of us and provide engaging experiments that students will find interesting 2 emphasize topics that students find particularly challenging in the general chemistry lecture course and 3 create a laboratory environment that encourages students on occasion to solve puzzles and not just follow recipes All too often students view general chemistry lab as a boring exercise in which an exact set of instructions is followed leading to an answer that in many cases results in a good grade regardless of how much learning has taken place To these students the successful lab is the one that takes the least amount of time Unfortunately a huge opportunity to get students truly turned on to science is missed To us the laboratory represents high stakes ground for engagement and relatively low stakes for grading as the laboratory is typically a single credit course or minor component to the lecture grade Thus while the rigor of the experiments in this manual can be tuned to meet the needs of the instructor our hope is that students will be encouraged to play safely with chemical concepts and laboratory techniques with grades simply being a natural consequence of their laboratory actions To facilitate such a mindset this manual has been written to provide instructors with a weekly tool that can attract and keep student interest while providing important connections to the material covered in an atoms first lecture course Our philosophy student curiosity leads to engagement which leads to discovery which leads to learning The manual is for a freshman level general chemistry laboratory course and serves as an ideal supplement for any atoms first general chemistry textbook such as *Chemistry Atoms First* by Julia Burdge and Jason Overby It is designed for students at all levels from those seeing chemistry for the first time to chemistry majors

A Laboratory Manual of Organic Chemistry for Beginners Arnold Frederick Holleman, 1913

Journal of Chemical Education, 1926 Includes Report of New England Association of Chemistry Teachers and Proceedings of the Pacific Southwest Association of Chemistry Teachers

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