

1. A solution is made by dissolving some salt in a beaker of water. The salt is referred to as the
A. solute B. filtrate C. solution D. solvent
2. When KCl dissolves in water, the following will be produced.
A. K and Cl B. K^+ and Cl^- C. K and Cl_2 D. K^+ and Cl_2
3. What is the concentration, in % (m/v) of a solution containing 15.0 g KCl in 600.0 mL solution?
A. 5.00% B. 2.60% C. 0.200% D. 2.50%
4. A 5.00 mL sample of solution has 2.8×10^{-4} g of calcium ions. The ppm concentration is
A. 18 ppm B. 56 ppm C. 2.8×10^{-5} ppm D. 2.8×10^7 ppm
5. How many mL of 2.50% (m/v) salt solution would contain 1.30 g of salt?
A. 105 mL B. 104 mL C. 45.0 mL D. 72.0 mL
6. What is the molarity of a solution that contains 1.50 mol HCl in 2.50 L of solution?
A. 1.67 M B. 0.600 M C. 1.20 M D. 1.40 M
7. How many moles of $C_{10}H_{22}O_{11}$ are needed to prepare 2.50 L of 0.300 M solution?
A. 0.750 mol B. 0.430 mol C. 8.33 mol D. 1.20 mol
8. What is the molarity of a $NaNO_3$ solution made by diluting 250.0 mL of a 1.60 M solution to a final volume of 400. mL?
A. 1.20 M B. 1.00 M C. 0.200 M D. 0.160 M
9. In the process known as osmosis, _____ moves through a semi-permeable membrane into an area of lower _____ concentration.
A. water, water B. solute, solute C. water, solute D. solute, water
10. What is the osmolarity of 0.800M $(NH_4)_3PO_4$ solution?
A. 3.20 osmol B. 2.40 osmol C. 1.60 osmol D. 8.00 osmol

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Table of Contents Chem Tutor 5 Self Check Activity Answers

1. Understanding the eBook Chem Tutor 5 Self Check Activity Answers
 - The Rise of Digital Reading Chem Tutor 5 Self Check Activity Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chem Tutor 5 Self Check Activity Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chem Tutor 5 Self Check Activity Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chem Tutor 5 Self Check Activity Answers
 - Personalized Recommendations
 - Chem Tutor 5 Self Check Activity Answers User Reviews and Ratings
 - Chem Tutor 5 Self Check Activity Answers and Bestseller Lists
5. Accessing Chem Tutor 5 Self Check Activity Answers Free and Paid eBooks
 - Chem Tutor 5 Self Check Activity Answers Public Domain eBooks
 - Chem Tutor 5 Self Check Activity Answers eBook Subscription Services
 - Chem Tutor 5 Self Check Activity Answers Budget-Friendly Options
6. Navigating Chem Tutor 5 Self Check Activity Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Chem Tutor 5 Self Check Activity Answers Compatibility with Devices
 - Chem Tutor 5 Self Check Activity Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Chem Tutor 5 Self Check Activity Answers
 - Highlighting and Note-Taking Chem Tutor 5 Self Check Activity Answers
 - Interactive Elements Chem Tutor 5 Self Check Activity Answers
8. Staying Engaged with Chem Tutor 5 Self Check Activity Answers
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chem Tutor 5 Self Check Activity Answers
9. Balancing eBooks and Physical Books Chem Tutor 5 Self Check Activity Answers
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chem Tutor 5 Self Check Activity Answers
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chem Tutor 5 Self Check Activity Answers
- Setting Reading Goals Chem Tutor 5 Self Check Activity Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chem Tutor 5 Self Check Activity Answers
- Fact-Checking eBook Content of Chem Tutor 5 Self Check Activity Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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