

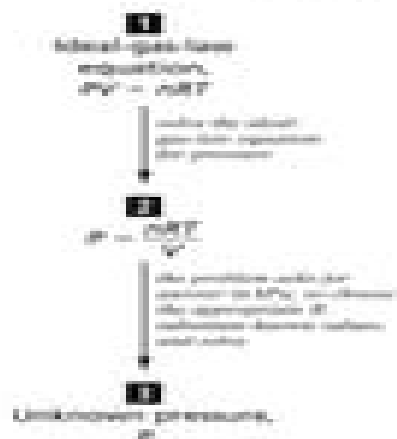
**CHEMFILE MINI-GUIDE TO PROBLEM SOLVING**

- *What are you asked to find?* the pressure of the gas in the cylinder

Items	Data
Amount of gas, $n$	5.00 mol
Volume of gas in cylinder, $V$	20.0 L
Temperature of gas, $t$	35°C
Kelvin temperature of gas, $T$	$(25 + 273) \text{ K} = 298 \text{ K}$
Ideal gas constant, $R$	$0.0821 \text{ L} \cdot \text{atm/mol} \cdot \text{K}$ or $8.314 \text{ L} \cdot \text{JPa/mol} \cdot \text{K}$
Pressure in cylinder, $P$	? kPa

**2. PLAN**

- *What steps are needed to calculate the new pressure of the gas?* Rearrange the ideal-gas-law equation to solve for  $P$ , substitute known quantities, and calculate.



$$PV = nRT$$

Solve the ideal-gas-law equation for  $P$ , the unknown quantity.

$$P = \frac{nRT}{V}$$

# Chemfile Mini Guide To Problem Solving Answer Key

**Rachel Turoscy**



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