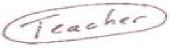
Stoichiometry III - Theoretical Yield



If we know what chemical reaction is occurring, we can use mole ratios to determine the limiting and excess reactants. We can do this by using the starting masses of each reactant to see how much of a product could be produced if all of the reactant was consumed. We then compare the number of moles produced by each reactant, and the smallest value is the theoretical yield - the amount of a product produced if all of the limiting reactant is consumed.

Practice Problems

- Given the following reaction: C₃H₈ +5O₅ →3CO₅ +4H₅O
 - a) Determine the limiting and excess reagents if you started with 14.8 g of C₂H₀ and 33.44 g of O₂.

- b) What is the theoretical yield of H₂O. 15-1a H₂O
- Given the following reaction: Al₂(SO₃)₃ + €NaOH → 3Na₂SO₃ + 2Al(OH)₃
 - a) Determine the limiting and excess reagents if 8.1 g of Al₂(SO₂), is reacted with 8.1 g of NaOH

- b) What is the theoretical yield of aluminum hydroxide. 4.3 A1 (0H)3
- 3. Given the following reaction: 4Al₂O₃ +9Fe →3Fe₃O₄ +8Al

a) Determine the limiting and excess reagents if 25.4 g of Al₂O₃ is reacted with 10.2 g of Fe.

b) What is the theoretical yield of iron(II,III) oxide (aka magnetite).

Chapter 3 Stoichiometry Exercises Answers

Theodore L. Brown, H. Eugene LeMay
Jr., Bruce E. Bursten, Catherine
Murphy, Patrick Woodward, Steven
Langford, Dalius Sagatys, Adrian
George

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CHEMISTRY SILBERBERG, 2003

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web announce the name of the winner of the contest highlight the efforts of the winner tell the readers what prize the winner will get congratulate the winner in a warm manner

30 congratulations to all the winners and participants - May 01 2022

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message hi amy congratulations you ve won the 500 wishpond gift card grand prize in our 500 summer giveaway contest to claim your prize please follow these steps confirm that you meet all of the entry

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