

KEY Review: Worksheet on Balancing Redox Equations

Two methods are often mentioned for balancing redox reactions: the half reaction method and the change in oxidation method. They actually involve the same procedure. In the first case you separate out the oxidation and reduction half reaction and in the second case, you do it all at once. I prefer the latter. The half reaction method is shown in the text but I will explain the change in oxidation method here.

Change in Oxidation Procedure:

- Write out as much of the unbalanced reaction as possible
- Assign oxidation numbers
- Draw brackets to connect the atoms that are oxidized and the atoms that are reduced. Write the net increases and decreases in electrons.
- Find the factors that create the least common multiple and use these to assign balanced stoichiometry for the reaction.

Example 1: No O or H atoms in equation

a. $\text{FeCl}_2 + \text{SnCl}_4 \rightarrow \text{SnCl}_2 + \text{FeCl}_3$ Write out unbalanced equation

b. $\overset{+2}{\text{Fe}}\overset{-1}{\text{Cl}}_2 + \overset{+4}{\text{Sn}}\overset{-1}{\text{Cl}}_4 \rightarrow \overset{2+}{\text{Sn}}\overset{-1}{\text{Cl}}_2 + \overset{+3}{\text{Fe}}\overset{-1}{\text{Cl}}_3$ Assign oxidation numbers

c. $\text{FeCl}_2 + \text{SnCl}_4 \rightarrow \text{SnCl}_2 + \text{FeCl}_3$ Draw brackets connecting redox atoms



d. $2\text{FeCl}_2 + 1\text{SnCl}_4 \rightarrow 1\text{SnCl}_2 + 2\text{FeCl}_3$ Find common factor, assign stoichiometry
Least common multiple: 2



Chapter 2 worksheet Redox

Arshad Iqbal

A red circular graphic with a gradient, appearing as a stylized arrow or a decorative element, located to the right of the author's name.

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