

Chapter 4 Atomic Structure

Section 4.1 Studying Atoms**(pages 100–106)***This section discusses the development of atomic models.***Reading Strategy** (page 100)

Summarizing. As you read, complete the table about atomic models. For more information on this Reading Strategy, see the **Reading and Study Skills** in the **Skills and Reference Handbook** at the end of your textbook.

Atomic Models		
Scientist	Evidence	Model
Democritus	Ratio of masses in compounds	Indivisible, solid spheres
Bohr	Deflected beam	Electrons revolve in fixed orbits around a positively charged mass of nucleus
Rutherford	Deflection of alpha particles passing through gold foil	Positive, dense nucleus

Ancient Greek Models of Atoms (page 100)

- Democritus named the smallest particles of matter atoms because they could not be divided.
- List the four elements that Aristotle included in his model of matter.
 - earth
 - fire
 - water
 - air

Dalton's Atomic Theory (page 101)

- Is the following sentence true or false? John Dalton gathered evidence for the existence of atoms by measuring the masses of elements that reacted to form compounds. false
- What theory did Dalton propose to explain why the elements in a compound always join in the same way? All elements are composed of atoms. In a particular compound, atoms of different elements always combine the same way. Compounds contain atoms of two or more elements.
- Circle the letters of the sentences that represent the main points of Dalton's theory of atoms.
 - All elements are composed of atoms.
 - In a particular compound, atoms of different elements always combine the same way.
 - All atoms have the same mass.
 - Compounds contain atoms of more than one element.

Chapter 4 Section 2 Review Answer Key Physical Science

Hussin A.Rothana



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