

Name :

Date : \_\_\_\_\_

## Trends in Periodic Table

Use the periodic table and your knowledge of periodic trends to answer the following questions.

1. Arrange the following group of elements in the order of increasing ionization energy.  
a) Be, Mg, Sr  
b) Bi, Cs, Ba  
c) Na, Al, S  
d) Co, Cl, K
2. Which element in each pair has the highest electronegativity value? Circle the correct answer.  
a) Cl, F  
b) C, N  
c) Mg, Ne  
d) As, Ca
3. Why do atoms get smaller as you move from left to right in the periodic table?
4. Circle the atom with the largest atomic radius and put a square on the element with the smallest atomic radius.

Explain why you made these choices.

5. Circle the element with the highest ionization energy and put a square around the element with the lowest ionization energy.

Explain why you made these choices.

6. Circle the element with the highest electronegativity and put a square around the element with the lowest electronegativity.

Explain why you made these choices.

# Chapter 6 Identifying Periodic Trends Worksheet Answers

**Mary Louise Turgeon**



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## **Table of Contents Chapter 6 Identifying Periodic Trends Worksheet Answers**

1. Understanding the eBook Chapter 6 Identifying Periodic Trends Worksheet Answers
  - The Rise of Digital Reading Chapter 6 Identifying Periodic Trends Worksheet Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 6 Identifying Periodic Trends Worksheet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 6 Identifying Periodic Trends Worksheet Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 6 Identifying Periodic Trends Worksheet Answers
  - Personalized Recommendations
  - Chapter 6 Identifying Periodic Trends Worksheet Answers User Reviews and Ratings
  - Chapter 6 Identifying Periodic Trends Worksheet Answers and Bestseller Lists

5. Accessing Chapter 6 Identifying Periodic Trends Worksheet Answers Free and Paid eBooks
  - Chapter 6 Identifying Periodic Trends Worksheet Answers Public Domain eBooks
  - Chapter 6 Identifying Periodic Trends Worksheet Answers eBook Subscription Services
  - Chapter 6 Identifying Periodic Trends Worksheet Answers Budget-Friendly Options
6. Navigating Chapter 6 Identifying Periodic Trends Worksheet Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 6 Identifying Periodic Trends Worksheet Answers Compatibility with Devices
  - Chapter 6 Identifying Periodic Trends Worksheet Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 6 Identifying Periodic Trends Worksheet Answers
  - Highlighting and Note-Taking Chapter 6 Identifying Periodic Trends Worksheet Answers
  - Interactive Elements Chapter 6 Identifying Periodic Trends Worksheet Answers
8. Staying Engaged with Chapter 6 Identifying Periodic Trends Worksheet Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 6 Identifying Periodic Trends Worksheet Answers
9. Balancing eBooks and Physical Books Chapter 6 Identifying Periodic Trends Worksheet Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 6 Identifying Periodic Trends Worksheet Answers
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Chapter 6 Identifying Periodic Trends Worksheet Answers
  - Setting Reading Goals Chapter 6 Identifying Periodic Trends Worksheet Answers
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chapter 6 Identifying Periodic Trends Worksheet Answers
  - Fact-Checking eBook Content of Chapter 6 Identifying Periodic Trends Worksheet Answers
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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