

7-1 Review and Reinforcement

Ionic Bonding

Complete each of the following sentences by filling in the appropriate word or phrase from the list below.

negative
positive
crisscross

noble gas
electrons
anion

octet
ions
cation

electron configuration
polyatomic
empirical formula

1. An ionic bond is an attraction between oppositely charged ions.
2. Anions have a negative charge.
3. An atom becomes an ion by losing or gaining electrons.
4. The octet rule states that atoms tend to gain, lose, or share electrons in order to acquire a full set of valence electrons.
5. When sodium and chlorine form an ionic bond, both ions acquire the electron configuration of a(n) noble gas.
6. The atoms that make up polyatomic ions are bonded together by covalent bonds.
7. The crisscross method can be used to write the formula for an ionic compound.
8. The empirical formula of a compound denotes the ratio of ions in the compound.

If the statement is true, write "true." If it is false, change the underlined word or words to make it true. Write your answer on the line.

- | | |
|------------------------|--|
| <u>losing</u> | 9. Calcium becomes a monatomic cation by <u>gaining</u> two electrons. |
| <u>true</u> | 10. A cation has a <u>positive</u> charge. |
| <u>true</u> | 11. A binary ionic compound contains only <u>one</u> kind of cation and one kind of anion. |
| <u>polyatomic</u> | 12. <u>Monatomic</u> ions consist of more than one atom. |
| <u>seven</u> | 13. The Lewis dot diagram for chlorine, a group 7A element, has <u>six</u> electrons. |
| <u>true (released)</u> | 14. A great deal of energy is <u>produced</u> when an ionic compound is formed. |

Chapter 5 Section 1 Reinforcement Answers

TD Snyder



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Table of Contents Chapter 5 Section 1 Reinforcement Answers

1. Understanding the eBook Chapter 5 Section 1 Reinforcement Answers
 - The Rise of Digital Reading Chapter 5 Section 1 Reinforcement Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 5 Section 1 Reinforcement Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 5 Section 1 Reinforcement Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 5 Section 1 Reinforcement Answers
 - Personalized Recommendations
 - Chapter 5 Section 1 Reinforcement Answers User Reviews and Ratings

- Chapter 5 Section 1 Reinforcement Answers and Bestseller Lists
- 5. Accessing Chapter 5 Section 1 Reinforcement Answers Free and Paid eBooks
 - Chapter 5 Section 1 Reinforcement Answers Public Domain eBooks
 - Chapter 5 Section 1 Reinforcement Answers eBook Subscription Services
 - Chapter 5 Section 1 Reinforcement Answers Budget-Friendly Options
- 6. Navigating Chapter 5 Section 1 Reinforcement Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 5 Section 1 Reinforcement Answers Compatibility with Devices
 - Chapter 5 Section 1 Reinforcement Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 5 Section 1 Reinforcement Answers
 - Highlighting and Note-Taking Chapter 5 Section 1 Reinforcement Answers
 - Interactive Elements Chapter 5 Section 1 Reinforcement Answers
- 8. Staying Engaged with Chapter 5 Section 1 Reinforcement Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 5 Section 1 Reinforcement Answers
- 9. Balancing eBooks and Physical Books Chapter 5 Section 1 Reinforcement Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 5 Section 1 Reinforcement Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 5 Section 1 Reinforcement Answers
 - Setting Reading Goals Chapter 5 Section 1 Reinforcement Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 5 Section 1 Reinforcement Answers
 - Fact-Checking eBook Content of Chapter 5 Section 1 Reinforcement Answers
 - Distinguishing Credible Sources

13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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