TYPES OF CHEMICAL REACTIONS

DISPLACEMENT (SINGLE AND DOUBLE)

INSTRUCTIONS:

Balance the equations below. Indicate the type of displacement reaction by writing SINGLE or DOUBLE in the space provided. Examples are provided below.

 $1Cu + 2 AgNO_3 \rightarrow 2 Ag + 1 Cu(NO_3)_2$

Reaction: SINGLE DISPLACEMENT

1 CuSO₄ + 2 NaOH → 1 Na₂SO₄ + 1 Cu(OH)₂

Reaction: DOUBLE DISPLACEMENT

QUESTIONS:

1)
$$_Cl_2 + _KI \rightarrow _KCI + _I_2$$
 _____DISPLACEMENT

2)
$$K_2CO_3 + BaCl_2 \rightarrow KCl + BaCO_3$$
 DISPLACEMENT

3)
$$Na + MgCl_2 \rightarrow NaCl + Mg$$
 DISPLACEMENT

4)
$$_Al + _CuCl_2 \rightarrow _AlCl_3 + _Cu$$
 _____DISPLACEMENT

5)
$$_Al + _Pb(NO_3)_2 \rightarrow _Al(NO_3)_3 + _Pb$$
 ____DISPLACEMENT

7)
$$_Al_2(SO_4)_3 + _Ca_3(PO_4)_2 \rightarrow _AlPO_4 + _CaSO_4$$
 DISPLACEMENT

8)
$$_FeBr_2 + _K_2CO_3 \rightarrow _FeCO_3 + _KBr$$
 DISPLACEMENT

10) __CuS + _ KCl
$$\rightarrow$$
 _ CuCl₂ + _ K₂S _ _ _ DISPLACEMENT

Double Displacement Reaction Answers

Morris Hein, Susan Arena

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