

Double Replacement Reactions and Solubility

Post-Lab Questions: Review your Observations Table. Complete AND balance each equation below. If your results yielded a precipitate, then write the net ionic equation in the box to the right. MAKE SURE TO SHOW STATES/PHASES for every reactant and product! Remember that all nitrates are soluble in water. If you did not see signs of a precipitate, then write NONE in the box to the right.

1) $\text{Na}_2\text{CO}_{3(aq)} + \text{Al}(\text{NO}_3)_3(aq) \rightarrow$	$6\text{NaNO}_3(aq) + \text{Al}(\text{OH})_3(s)$	None
2) $\text{NaCl}_{(aq)} + \text{NH}_4\text{NO}_3(aq) \rightarrow$	$\text{NaNO}_3(aq) + \text{NH}_4\text{Cl}(aq)$	None
3) $\text{NaI}_{(aq)} + \text{AgNO}_{3(aq)} \rightarrow$	$\text{NaNO}_3(aq) + \text{AgI}(s)$	$\text{Ag}^+(aq) + \text{I}^-(aq) \rightarrow \text{AgI}(s)$
4) $\text{Na}_3\text{PO}_{4(aq)} + \text{Cu}(\text{NO}_3)_2(aq) \rightarrow$	$6\text{NaNO}_3(aq) + (\text{Cu}_3(\text{PO}_4)_2)(s)$	$3\text{Cu}^{2+}(aq) + 2\text{PO}_4^{3-}(aq) \rightarrow \text{Cu}_3(\text{PO}_4)_2(s)$

Use your observations table to answer the following. *ALL ions must include a symbol AND a charge or it is wrong!*

5 a) The symbols and charges of the anions that produced the fewest (<3) precipitates are _____

b) Salts of these anions may be described as SOLUBLE except for those with the cations _____

6a) The symbols and charges of the anion that produced the most (>5) precipitates are _____

b) Salts of these anions may be described as INSOLUBLE except for those with the cations _____

7. Although you did not test the bromide ion (Br^-), would you expect most bromide salts (AlBr_3 , BaBr_2 , CaBr_2 , etc) to be soluble or insoluble? Hint: Locate bromine on the periodic table. _____

8. Explain your answer in #7 _____

9. What cation combined with bromide is an exception to the rule in #7? _____

10-14 Summarize the general solubility rules of ionic compounds in the table below. Two have been done for you.

Ion	Generally soluble or insoluble?	Exceptions
9) carbonates		
10) chlorides, bromides, iodides		
11) hydroxides		
nitrates	Soluble	None
12) phosphates		

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G Thomas



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Introductory Chemistry Steve Russo, Mike Silver, 2002 Steve Russo and Mike Silver turn chemistry into a memorable story that engages readers and provides the context they need to understand and remember core concepts The book builds interesting applications and well designed illustrations into the narrative to get and hold attention then builds confidence with integrated active learning activities Readers make the connections between concepts and the problem solving techniques they need to master as they read The new edition strengthens this conceptual approach and presents additional quantitative techniques in key areas Readers will find enhanced support for quantitative problem solving and more challenging questions at the end of each chapter in addition to the wealth of technology based support on The Chemistry Place tm Special Edition and on The Chemistry of Life CD ROM For college instructors and students Essentials of Introductory Chemistry Russo Steve Silver Michael, Steve Russo, 2001-12 Introductory Chemistry Third Edition helps readers master the quantitative skills and conceptual understanding they need to gain a deep understanding of chemistry Unlike other books on the market that emphasize rote memory of problem solving algorithms Introductory Chemistry takes a conceptual approach with the idea that focusing on the concepts behind chemical equations helps readers become more proficient problem solvers What Is Chemistry The Numerical Side of Chemistry The Evolution of Atomic Theory The Modern Model of the Atom 1 Chemical Bonding and Nomenclature The Shape of Molecules Chemical Reactions Stoichiometry and the Mole The Transfer of Electrons from One Atom to Another in a Chemical Reaction Intermolecular Forces and the Phases of Matter What If There Were No Intermolecular Forces The Ideal Gas Solutions When Reactants Turn into Products Chemical Equilibrium Electrolytes Acids and Bases For all readers interested in introductory chemistry **Ei Engineering Conference Index: pt. 1. Civil, environmental, and geological engineering**, 1985 **Stereoselective Synthesis of Drugs and Natural Products** Vasyi Andrushko, Natalia Andrushko, 2013-08-16 Brings together the best tested and proven stereoselective synthetic methods Both the chemical and pharmaceutical industries are increasingly dependent on stereoselective synthetic methods and strategies for the generation of new chiral drugs and natural products that offer specific 3 D structures With the publication of Stereoselective Synthesis of Drugs and Natural Products researchers can turn to this comprehensive two volume work to guide them through all the core methods for the synthesis of chiral drugs and natural products Stereoselective Synthesis of Drugs and Natural Products features contributions from an international team

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