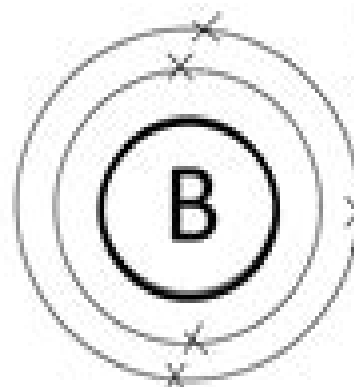


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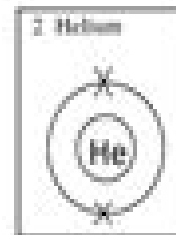
1 Hydrogen



The X is an electron.

The circle O is a shell that electrons sit on.

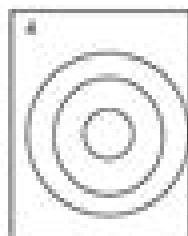
2 Helium



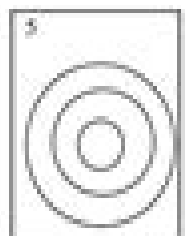
3 Lithium



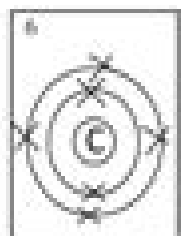
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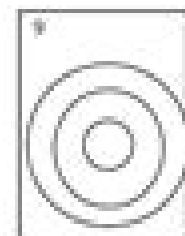
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10



11 Sodium



12



13



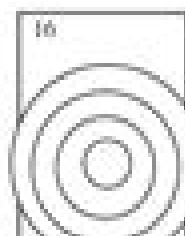
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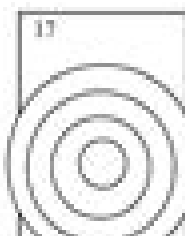
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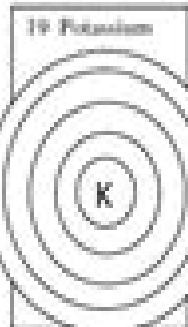
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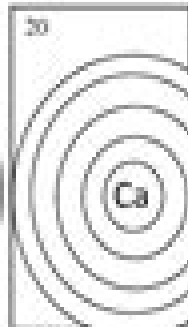
18



19 Potassium



20



Fill in each box with its electron shell pattern.

**Remember:**

- 1) The 1<sup>st</sup> shell can have up to **2** electrons
- 2) The 2<sup>nd</sup> shell can have up to **8** electrons
- 3) The 3<sup>rd</sup> shell can have up to **8** electrons

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**Manufacturing Processes** H. C. Kazanas, Glenn E. Baker, Thomas Gregor, 1981

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web elements in the second row of the periodic table place their electrons in the  $2n$  shell as well as the  $1n$  shell for instance lithium li has three electrons two fill the  $1s$  orbital and the third is placed in the  $2s$  orbital giving an electron configuration of  $1s^2 2s^1$

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web the electron removed during the ionization of beryllium he  $2s^2$  is an  $s$  electron whereas the electron removed during the ionization of boron he  $2s^2 2p^1$  is a  $p$  electron this results in a lower first ionization energy for boron even though its nuclear charge is greater by one proton

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web predictions from reliable sources have been used for these elements grayed out electron numbers indicate subshells filled to their maximum bracketed noble gas symbols on the left represent inner configurations that are the same in each period written out these are he  $2$  helium  $1s^2$  ne  $10$  neon  $1s^2 2s^2 2p^6$

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