

Reference Tables for Physical Setting/EARTH SCIENCE

Radioactive Decay Data

RADIOACTIVE ISOTOPE	DISINTEGRATION	HALF-LIFE (years)
Carbon-14	$^{14}\text{C} \rightarrow ^{14}\text{N}$	5.7×10^3
Potassium-40	$^{40}\text{K} \rightarrow ^{40}\text{Ar}$ $^{40}\text{K} \rightarrow ^{40}\text{Ca}$	1.3×10^9
Uranium-238	$^{238}\text{U} \rightarrow ^{206}\text{Pb}$	4.5×10^9
Rubidium-87	$^{87}\text{Rb} \rightarrow ^{87}\text{Sr}$	4.9×10^{10}

Specific Heats of Common Materials

MATERIAL	SPECIFIC HEAT (Joules/gram • °C)
Liquid water	4.18
Solid water (ice)	2.11
Water vapor	2.00
Dry air	1.01
Basalt	0.84
Granite	0.79
Iron	0.45
Copper	0.38
Lead	0.13

Equations

$$\text{Eccentricity} = \frac{\text{distance between foci}}{\text{length of major axis}}$$

$$\text{Gradient} = \frac{\text{change in field value}}{\text{distance}}$$

$$\text{Rate of change} = \frac{\text{change in value}}{\text{time}}$$

$$\text{Density} = \frac{\text{mass}}{\text{volume}}$$

Properties of Water

Heat energy gained during melting	334 J/g
Heat energy released during freezing	334 J/g
Heat energy gained during vaporization	2260 J/g
Heat energy released during condensation	2260 J/g
Density at 3.98°C	1.0 g/mL

Average Chemical Composition of Earth's Crust, Hydrosphere, and Troposphere

ELEMENT (symbol)	CRUST		HYDROSPHERE	TROPOSPHERE
	Percent by mass	Percent by volume	Percent by volume	Percent by volume
Oxygen (O)	46.10	94.04	33.0	21.0
Silicon (Si)	28.20	0.88		
Aluminum (Al)	8.23	0.48		
Iron (Fe)	5.63	0.49		
Calcium (Ca)	4.15	1.18		
Sodium (Na)	2.36	1.11		
Magnesium (Mg)	2.33	0.33		
Potassium (K)	2.09	1.42		
Nitrogen (N)				78.0
Hydrogen (H)			66.0	
Other	0.91	0.07	1.0	1.0

2011 EDITION

This edition of the Earth Science Reference Tables should be used in the classroom beginning in the 2011–12 school year. The first examination for which these tables will be used is the January 2012 Regents Examination in Physical Setting/Earth Science.

Eurypterus remipes

New York State Fossil



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J Rink



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Table of Contents Earth Science Reference Tables Workbook 3rd Edition Answers

1. Understanding the eBook Earth Science Reference Tables Workbook 3rd Edition Answers
 - The Rise of Digital Reading Earth Science Reference Tables Workbook 3rd Edition Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Earth Science Reference Tables Workbook 3rd Edition Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Earth Science Reference Tables Workbook 3rd Edition Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Earth Science Reference Tables Workbook 3rd Edition Answers
 - Personalized Recommendations
 - Earth Science Reference Tables Workbook 3rd Edition Answers User Reviews and Ratings
 - Earth Science Reference Tables Workbook 3rd Edition Answers and Bestseller Lists
5. Accessing Earth Science Reference Tables Workbook 3rd Edition Answers Free and Paid eBooks
 - Earth Science Reference Tables Workbook 3rd Edition Answers Public Domain eBooks
 - Earth Science Reference Tables Workbook 3rd Edition Answers eBook Subscription Services
 - Earth Science Reference Tables Workbook 3rd Edition Answers Budget-Friendly Options
6. Navigating Earth Science Reference Tables Workbook 3rd Edition Answers eBook Formats

- ePub, PDF, MOBI, and More
 - Earth Science Reference Tables Workbook 3rd Edition Answers Compatibility with Devices
 - Earth Science Reference Tables Workbook 3rd Edition Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Earth Science Reference Tables Workbook 3rd Edition Answers
 - Highlighting and Note-Taking Earth Science Reference Tables Workbook 3rd Edition Answers
 - Interactive Elements Earth Science Reference Tables Workbook 3rd Edition Answers
 8. Staying Engaged with Earth Science Reference Tables Workbook 3rd Edition Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Earth Science Reference Tables Workbook 3rd Edition Answers
 9. Balancing eBooks and Physical Books Earth Science Reference Tables Workbook 3rd Edition Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Earth Science Reference Tables Workbook 3rd Edition Answers
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Earth Science Reference Tables Workbook 3rd Edition Answers
 - Setting Reading Goals Earth Science Reference Tables Workbook 3rd Edition Answers
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Earth Science Reference Tables Workbook 3rd Edition Answers
 - Fact-Checking eBook Content of Earth Science Reference Tables Workbook 3rd Edition Answers
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
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 - Interactive and Gamified eBooks

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