

The Cell: Transport Mechanisms and Cell Permeability – Wet Lab

Objectives

1. To *define differential permeability* and explain the difference between *active* and *passive* processes of cellular transport.
2. To define diffusion (*simple diffusion* and *osmosis*); *hypotonic*, and *hypertonic* solutions; *active transport*; *vesicular transport*; and *exocytosis*, *phagocytosis*, and *pinocytosis*.
3. To describe the processes that account for the movement of substances across the plasma membrane and to indicate the driving force for each.
4. To determine which way substances will move passively through a differentially permeable membrane (given appropriate information on concentration differences).

Because of its molecular composition, the plasma membrane is selective about what passes through it. It allows nutrients to enter the cell but keeps out undesirable substances. By the same token, valuable cell proteins and other substances are kept within the cell, and wastes pass to the exterior. This property is known as **differential**, or **selective, permeability**. Transport through the plasma membrane occurs in two basic ways. In **passive processes**, concentration or pressure differences drive the movement. In **active processes**, the cell provides energy (ATP) to power the transport process.

Passive Processes

The two important passive processes of membrane transport are *diffusion* and *filtration*. Diffusion is an important transport process for every cell in the body. By contrast, filtration usually occurs only across capillary walls. Only diffusion will be considered here.

Recall that all molecules possess *kinetic energy* and are in constant motion. At a specific temperature, given molecules have about the same average kinetic energy. Since kinetic energy is directly related to both mass and velocity ($KE = \frac{1}{2}mv^2$), smaller molecules tend to move faster. As molecules move about randomly at high speeds, they collide and ricochet off one another, changing direction with each collision.

Although individual molecules cannot be seen, the random motion of small particles suspended in water can be observed. This is called **Brownian movement**.

Diffusion

When a **concentration gradient** (difference in concentration) exists, the net effect of this random molecular movement is that the molecules eventually become evenly distributed throughout the environment, that is, the process called diffusion occurs. Hence, **diffusion** is the movement of molecules from a region of their higher concentration to a region of their lower concentration. Its driving force is the kinetic energy of the molecules themselves.

There are many examples of diffusion in nonliving systems. For example, if a bottle of ether was uncorked at the front of the laboratory, very shortly thereafter you would be nodding as the ether

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Timothy Stabler



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and synthetic life An appendix provides a Python toolkit for implementing ACs **A Novel Class of in Silico Devices to Study Drug Transport Through Intestinal Barriers** Lan Garmire,2007 **Cumulated Index Medicus** ,1983

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