

# Electrical Conductivity of Aqueous Solutions

## Objectives

The objectives of this laboratory are:

- To observe electrical conductivity of substances in various aqueous solutions
- To determine if the solution is a strong or weak electrolyte
- To interpret a chemical reaction by observing aqueous solution conductivity.

## Background

Electrical conductivity is based on the flow of electrons. Metals are good conductors of electricity because they allow electrons to flow through the entire piece of material. Thus, electrons flow like a "sea of electrons" through metals. In comparison, distilled water is a very poor conductor of electricity since very little electricity flows through water. Highly ionized substances are strong electrolytes. Strong acids and salts are strong electrolytes because they completely ionize (dissociate or separate) in solution. The ions carry the electric charge through the solution thus creating an electric current. When connected to a power source, a strong electrolyte solution can conduct enough electricity to light a bulb, as shown in the figure below.



Slightly ionized substances are weak electrolytes. Weak acids and bases would be categorized as weak electrolytes because they do not completely dissociate in solution. Weak electrolyte solutions would cause the bulb to light, but not as brightly.

Substances that do not conduct an electric current are called non-electrolytes. Non-electrolytes do not ionize; they do not contain moveable ions. The lightbulb will not turn on because there are no ions to carry the electric current. The table below lists examples of strong, weak and non-electrolytes.

Strong Electrolytes	Weak Electrolytes	Non-Electrolytes
Strong Acids (e.g. HCl)	Weak Acids (e.g. HCl)	Molecular compounds (e.g. C <sub>6</sub> H <sub>12</sub> O <sub>6</sub> )
Strong Bases (e.g. NaOH)	Weak Bases (e.g. NaOH)	
Soluble Salts (e.g. NaCl)	Slightly Soluble Salts (e.g. NaCl)	

# Conductivity Of Aqueous Solutions Lab Answers

**L Darling-Hammond**



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