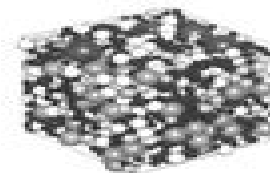




Key

Density Practice Problems



The density of a substance is a measure of how much mass is packed into a certain volume of the substance. Substances with a high density, like steel, have molecules that are packed together tightly. Substances with a low density, like cork, have fewer molecules packed into the same amount of space.

The density of a substance can be found by dividing its mass by its volume. As long as a substance is homogeneous, the size or shape of the sample doesn't matter. The density will always be the same. This means that a steel paper clip has the same density as a steel girder used to build a bridge.

$$\text{Density} = \text{Mass} \div \text{Volume}$$

$$D = \frac{m}{V}$$

Use the density formula to solve the following problems. Show all work and the answer must have the correct units. Remember that volume can have different forms. A block of ice with a volume of 3 cm³ would be 3 mL of liquid after being melted.

1. What is the density of CO gas if 0.196 g occupies a volume of 100 ml?

$$m = 0.196 \text{ g}$$

$$V = 100 \text{ ml}$$

$$D = \frac{0.196 \text{ g}}{100 \text{ ml}}$$

Answer: $0.00196 \text{ g/ml} = 0.002 \text{ g/ml}$

2. A block of wood 3 cm on each side has a mass of 27 g. What is the density of the block? (Hint, don't forget to find the volume of the wood.)

$$V = L \times W \times H$$

$$V = 3 \times 3 \times 3$$

$$V = 27 \text{ cm}^3$$

$$D = \frac{27 \text{ g}}{27 \text{ cm}^3}$$

$$D = 1 \text{ g/cm}^3$$

$$m = 27 \text{ g}$$

Answer: 1 g/cm^3

3. A student measures a rectangular block of an unknown substance. The block measures 12.1 cm by 17.5 cm by 5.8 cm. The student weighs the block and determines its mass to be 4250.112 g. What is the density of this material? Watch the sig figs.

$$V = L \times W \times H$$

$$V = 12.1 \times 17.5 \times (5.8) \quad 2 \text{ s.f.}$$

$$V = 1228.15 \text{ cm}^3$$

$$D = \frac{4250.112 \text{ g}}{1228.15 \text{ cm}^3}$$

Answer: $3.46 \text{ g/cm}^3 = 3.5 \text{ g/cm}^3$

4. An irregularly shaped stone was lowered into a graduated cylinder holding a volume of water equal to 3.5 ml. The height of the water rose to 8.0 ml. If the mass of the stone was 27 g, what was its density?

$$V = 8.0 \text{ mL} - 3.5 \text{ mL}$$

$$m = 27 \text{ g}$$

$$V = 4.5 \text{ mL}$$

$$D = \frac{27 \text{ g}}{4.5 \text{ mL}} = 6$$

Answer: 6.0 g/mL

Density Practice Problem Answer Key Chapter 7

R Pring



Density Practice Problem Answer Key Chapter 7:

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Table of Contents Density Practice Problem Answer Key Chapter 7

1. Understanding the eBook Density Practice Problem Answer Key Chapter 7
 - The Rise of Digital Reading Density Practice Problem Answer Key Chapter 7
 - Advantages of eBooks Over Traditional Books
2. Identifying Density Practice Problem Answer Key Chapter 7
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Density Practice Problem Answer Key Chapter 7
 - User-Friendly Interface
4. Exploring eBook Recommendations from Density Practice Problem Answer Key Chapter 7
 - Personalized Recommendations
 - Density Practice Problem Answer Key Chapter 7 User Reviews and Ratings
 - Density Practice Problem Answer Key Chapter 7 and Bestseller Lists
5. Accessing Density Practice Problem Answer Key Chapter 7 Free and Paid eBooks
 - Density Practice Problem Answer Key Chapter 7 Public Domain eBooks
 - Density Practice Problem Answer Key Chapter 7 eBook Subscription Services
 - Density Practice Problem Answer Key Chapter 7 Budget-Friendly Options
6. Navigating Density Practice Problem Answer Key Chapter 7 eBook Formats

- ePub, PDF, MOBI, and More
 - Density Practice Problem Answer Key Chapter 7 Compatibility with Devices
 - Density Practice Problem Answer Key Chapter 7 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Density Practice Problem Answer Key Chapter 7
 - Highlighting and Note-Taking Density Practice Problem Answer Key Chapter 7
 - Interactive Elements Density Practice Problem Answer Key Chapter 7
 8. Staying Engaged with Density Practice Problem Answer Key Chapter 7
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Density Practice Problem Answer Key Chapter 7
 9. Balancing eBooks and Physical Books Density Practice Problem Answer Key Chapter 7
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Density Practice Problem Answer Key Chapter 7
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Density Practice Problem Answer Key Chapter 7
 - Setting Reading Goals Density Practice Problem Answer Key Chapter 7
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Density Practice Problem Answer Key Chapter 7
 - Fact-Checking eBook Content of Density Practice Problem Answer Key Chapter 7
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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