

Stoichiometry Worksheet 2: Percent Yield

For each of the problems below:

- Write the balanced chemical equation
- Identify the given (with units) and what you want to find (with units)
- Show set up with units. Check sig figs, give final answer with units and label.

1. Using the Hoffman apparatus for electrolysis, a chemist decomposes 36 g of water into its gaseous elements. How many grams of hydrogen gas should she get (theoretical yield)?



Before $\underline{2 \text{ mol}}$ $\underline{0 \text{ mol}}$ $\underline{0 \text{ mol}}$

Change $\underline{-2 \text{ mol}}$ $\underline{+2 \text{ mol}}$ _____

After $\underline{0 \text{ mol}}$ $\underline{2 \text{ mol}}$

$$36 \text{ g H}_2\text{O} \times \frac{1 \text{ mol H}_2\text{O}}{18.0 \text{ g H}_2\text{O}} = 2.0 \text{ mol H}_2\text{O} \times \frac{2 \text{ mol H}_2}{2 \text{ mol H}_2\text{O}} = 2 \text{ mol H}_2 \times \frac{2.0 \text{ g H}_2}{1 \text{ mol H}_2} = 4.0 \text{ g H}_2$$

2. Recall that liquid sodium reacts with chlorine gas to produce sodium chloride. You want to produce 581 g of sodium chloride. How many grams of sodium are needed?



Before $\underline{9.94 \text{ mol}}$ $\underline{\text{xs mol}}$ $\underline{0 \text{ mol}}$

Change $\underline{-9.94 \text{ mol}}$ _____ $\underline{+9.94 \text{ mol}}$

After $\underline{0 \text{ mol}}$ $\underline{9.94 \text{ mol}}$

$$581 \text{ g NaCl} \times \frac{1 \text{ mol NaCl}}{58.5 \text{ g NaCl}} = 9.93 \text{ mol NaCl} \times \frac{2 \text{ mol Na}}{2 \text{ mol NaCl}} = 9.93 \text{ mol Na} \times \frac{23.0 \text{ g Na}}{1 \text{ mol Na}} = 228 \text{ g Na}$$

3. You eat 180.0 g of glucose (90 M&Ms). If glucose, $\text{C}_6\text{H}_{12}\text{O}_6$, reacts with oxygen gas to produce carbon dioxide and water, how many grams of oxygen will you have to breathe in to burn the glucose?



Before $\underline{1.0 \text{ mol}}$ $\underline{\text{xs mol}}$ $\underline{0 \text{ mol}}$ $\underline{0 \text{ mol}}$

Change $\underline{-1.0 \text{ mol}}$ _____ $\underline{-1.0 \text{ mol}}$

After $\underline{0 \text{ mol}}$ $\underline{\text{xs mol}}$

$$180.0 \text{ g C}_6\text{H}_{12}\text{O}_6 \times \frac{1 \text{ mol C}_6\text{H}_{12}\text{O}_6}{180 \text{ g C}_6\text{H}_{12}\text{O}_6} = 1.00 \text{ mol C}_6\text{H}_{12}\text{O}_6 \times \frac{6 \text{ mol O}_2}{1 \text{ mol C}_6\text{H}_{12}\text{O}_6} = 6.00 \text{ mol O}_2$$

$$6.0 \text{ mol O}_2 \times \frac{32.0 \text{ g O}_2}{1 \text{ mol O}_2} = 192.0 \text{ g O}_2$$

Chemistry Supplemental Stoichiometry Problems

Answer Key

Timo Ovaska



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1. Understanding the eBook Chemistry Supplemental Stoichiometry Problems Answer Key
 - The Rise of Digital Reading Chemistry Supplemental Stoichiometry Problems Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Supplemental Stoichiometry Problems Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Supplemental Stoichiometry Problems Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Supplemental Stoichiometry Problems Answer Key
 - Personalized Recommendations
 - Chemistry Supplemental Stoichiometry Problems Answer Key User Reviews and Ratings
 - Chemistry Supplemental Stoichiometry Problems Answer Key and Bestseller Lists
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 - Chemistry Supplemental Stoichiometry Problems Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Supplemental Stoichiometry Problems Answer Key
 - Highlighting and Note-Taking Chemistry Supplemental Stoichiometry Problems Answer Key
 - Interactive Elements Chemistry Supplemental Stoichiometry Problems Answer Key
- 8. Staying Engaged with Chemistry Supplemental Stoichiometry Problems Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Supplemental Stoichiometry Problems Answer Key
- 9. Balancing eBooks and Physical Books Chemistry Supplemental Stoichiometry Problems Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Supplemental Stoichiometry Problems Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chemistry Supplemental Stoichiometry Problems Answer Key
 - Setting Reading Goals Chemistry Supplemental Stoichiometry Problems Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chemistry Supplemental Stoichiometry Problems Answer Key
 - Fact-Checking eBook Content of Chemistry Supplemental Stoichiometry Problems Answer Key
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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