

Unit 2 Worksheet 3 – PVTn Problems

On each of the problems below, start with the given P, V, T, or n; then make a decision as to how a change in P, V, T, or n will affect the starting quantity, and then multiply by the appropriate factor. Draw particle diagrams of the initial and final conditions.

1. What would be the new pressure if 250 cm³ of gas at standard pressure is compressed to a volume of 150 cm³? (n and T = constant)

	P	T	V	n
Initial	1 atm	—	250 cm ³	—
Final	?	—	150 cm ³	—
Effect	↑	—	↓	—

$$P_2 = 1 \text{ atm} \times \frac{250 \text{ cm}^3}{150 \text{ cm}^3} = 1.7 \text{ atm}$$

or 1.340 atm / kg
or 1.704 Pa
or 25 psi

2. The pressure in a bicycle tire is 105 psi at 25°C in Fresno. You take the bicycle up to Huntington, where the temperature is – 5°C. What is the pressure in the tire? (V and n = constant)

	P	T	V	n
Initial	105 psi	298 K	—	—
Final	?	268 K	—	—
Effect	↓	↓	—	—

$$P_2 = 105 \text{ psi} \times \frac{268 \text{ K}}{298 \text{ K}} = 94.4 \text{ psi}$$

3. A sample of gas occupies 150 mL at 25 °C. What is its volume when the temperature is increased to 50°C? (P and n = constant)

	P	T	V	n
Initial	—	298 K	150 mL	—
Final	—	323 K	?	—
Effect	—	↑	↑	—

$$V_2 = 150 \text{ mL} \times \frac{323 \text{ K}}{298 \text{ K}} = 163 \text{ mL} \rightarrow 160 \text{ mL}$$

Chemistry Unit 2 Pvtn Problems Answers

F Rizvi



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Table of Contents Chemistry Unit 2 PvtN Problems Answers

1. Understanding the eBook Chemistry Unit 2 PvtN Problems Answers
 - The Rise of Digital Reading Chemistry Unit 2 PvtN Problems Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chemistry Unit 2 PvtN Problems Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chemistry Unit 2 PvtN Problems Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chemistry Unit 2 PvtN Problems Answers
 - Personalized Recommendations
 - Chemistry Unit 2 PvtN Problems Answers User Reviews and Ratings
 - Chemistry Unit 2 PvtN Problems Answers and Bestseller Lists
5. Accessing Chemistry Unit 2 PvtN Problems Answers Free and Paid eBooks
 - Chemistry Unit 2 PvtN Problems Answers Public Domain eBooks
 - Chemistry Unit 2 PvtN Problems Answers eBook Subscription Services
 - Chemistry Unit 2 PvtN Problems Answers Budget-Friendly Options

6. Navigating Chemistry Unit 2 Pvtn Problems Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chemistry Unit 2 Pvtn Problems Answers Compatibility with Devices
 - Chemistry Unit 2 Pvtn Problems Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chemistry Unit 2 Pvtn Problems Answers
 - Highlighting and Note-Taking Chemistry Unit 2 Pvtn Problems Answers
 - Interactive Elements Chemistry Unit 2 Pvtn Problems Answers
8. Staying Engaged with Chemistry Unit 2 Pvtn Problems Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chemistry Unit 2 Pvtn Problems Answers
9. Balancing eBooks and Physical Books Chemistry Unit 2 Pvtn Problems Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chemistry Unit 2 Pvtn Problems Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chemistry Unit 2 Pvtn Problems Answers
 - Setting Reading Goals Chemistry Unit 2 Pvtn Problems Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chemistry Unit 2 Pvtn Problems Answers
 - Fact-Checking eBook Content of Chemistry Unit 2 Pvtn Problems Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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