

Unit 5

Systems of Linear Equations and Inequalities 8 Lessons & Homework Sets

COMMON CORE ALGEBRA I



Common Core Algebra Emathinstruction Keys

**Laura Michelle Ferguson, Jessica Paige
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web start with 8.3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8.3 \text{ mole } \text{H}_2\text{S} \times \frac{2 \text{ mole } \text{H}_2\text{O}}{2 \text{ mole } \text{H}_2\text{S}} \times 18 \text{ g } \text{H}_2\text{O} = 149.4 \text{ g } \text{H}_2\text{O}$ is your theoretical yield so percent yield $\frac{137.1 \text{ g } \text{H}_2\text{O}}{149.4 \text{ g } \text{H}_2\text{O}} \times 100 = 91.77\%$ problems 6 8 are similar to problem 5

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web a write a balanced equation for the reaction $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$ b determine the theoretical yield of Al_2O_3 160.0 g Al $\times \frac{1 \text{ mol Al}}{2 \text{ mol Al}_2\text{O}_3} \times 101.96 \text{ g } \text{Al}_2\text{O}_3 = 81.2 \text{ g } \text{Al}_2\text{O}_3$ c determine the percent yield

yield actual x 100 260 0 g x 100 86 01 theoretical 302 3 g

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 $80 \text{ g } \text{I}_2\text{O}_5 \times \frac{1 \text{ mol } \text{I}_2\text{O}_5}{1 \text{ mol } \text{I}_2} \times \frac{1}{333.8 \text{ g } \text{I}_2\text{O}_5} \times 28 \text{ g } \text{CO} \times \frac{1 \text{ mol } \text{CO}}{28 \text{ g } \text{CO}} = 1 \text{ mol } \text{I}_2$

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