



# COVALENT BONDING

## Practice Problems

Do your answers to these very practice problems.

### SECTION 8.1 MOLECULAR COMPOUNDS

1. Draw a molecular formula for each of the following compounds.

- a. Silicon tetrachloride      b. Phosphorus pentachloride      c. Nitrogen dioxide  
d. Sulfur hexafluoride      e. Boron trifluoride

2. Write the formula for the following molecular compounds.

- a. Carbon disulfide      b. Dinitrogen pentoxide      c. Carbon monoxide  
d. Nitrogen dioxide      e. Sulfur dioxide

3. Draw the Lewis structure for each of the following molecular compounds.

a. Dichloromethane      b. Carbon tetrachloride      c. Phosphorus pentachloride      d. Nitrogen dioxide      e. Sulfur dioxide

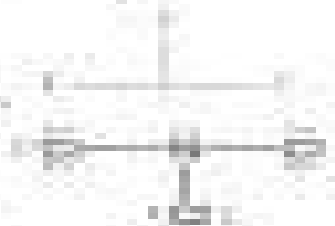
4. Draw the Lewis structure for each of the following molecular compounds. Label each atom as being an electrophile or nucleophile.

### SECTION 8.2 THE NATURE OF COVALENT BONDING

1. Draw the Lewis structure for hydrogen fluoride, HF.



2. Draw the Lewis structure for methane, CH<sub>4</sub>.



3. Draw the Lewis structure for carbon dioxide, CO<sub>2</sub>.



4. Draw the Lewis structure for acetylene, C<sub>2</sub>H<sub>2</sub>.

5. Draw the Lewis structure for each of the following compounds. Label each atom as being an electrophile or nucleophile.



6. Draw the Lewis structure for each of the following compounds.

### SECTION 8.3 BONDING: THE ORBITAL

1. Draw the Lewis structure for each of the following compounds. Label each atom as being an electrophile or nucleophile.

2. Draw the Lewis structure for each of the following compounds. Label each atom as being an electrophile or nucleophile.

# Covalent Bonding Practice Problems Chapter 8

**Ying-Ying Zheng**



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