

Name _____

Section _____

Experiment 3**Data and Calculations: Determination of the Equilibrium Constant for a Chemical Reaction**

Trials	Volume of 0.1M Fe^{3+} (mL)	Volume of 0.1M SCN^{-} (mL)	Volume of 0.1M HCl (mL)	Initial $[\text{Fe}^{3+}]$ (M)	Initial $[\text{SCN}^{-}]$ (M)		Equilibrium $[\text{FeSCN}^{2+}]$ (M)
	Initial	Initial	Initial	Initial	Initial	Initial	
1	5.00	1.00	2.00	0.010	0.002	0.000	0.000
2	5.00	1.00	3.00	0.010	0.002	0.000	0.000
3	5.00	1.00	2.00	0.010	0.002	0.000	0.000
4	5.00	1.00	1.00	0.010	0.002	0.000	0.000
5	5.00	1.00	0.00	0.010	0.002	0.000	0.000

If Method B has been used, Section 3.1.1.1 of the manual should be used to determine the equilibrium constant.

6. Calculate the equilibrium constant for the reaction.

7. Calculate the standard free energy change for the reaction.

Processing the Data

A. Calculation of K_c assuming the reaction:

$\text{Fe}^{3+} + \text{SCN}^{-} \rightleftharpoons \text{FeSCN}^{2+}$

This calculation is more easily done by knowing how to change K_c to the standard free energy change of the reaction. If you are using Method B, use the values of K_c from the previous section.

Step 1: Find the initial number of moles of Fe^{3+} and SCN^{-} in the reaction. (Use the initial concentrations of Fe^{3+} and SCN^{-} from the previous section.)


Step 2: From the experimentally determined value of K_c (from the previous section) and the initial concentrations of Fe^{3+} and SCN^{-} , calculate the equilibrium concentrations of Fe^{3+} , SCN^{-} , and FeSCN^{2+} . (Use the initial concentrations of Fe^{3+} and SCN^{-} from the previous section.)

Calculate the equilibrium constant.

Calculate the standard free energy change for the reaction.

Calculations With A Chemical Reaction Lab Answers

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Hassan,Prathamesh Churi**



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