

- 13 These calculations have used $L = 6.02 \times 10^{23}$
 (a) 7.22×10^{22} (b) 3.01×10^{23}
 (c) 1.20×10^{23}

14 0.53 mol H

15 0.250 mol

- 16 (a) 262.87 g mol⁻¹ (b) 176.14 g mol⁻¹
 (c) 164.10 g mol⁻¹ (d) 248.22 g mol⁻¹

17 189.1 g

18 1.5 mol

19 0.0074 mol Cl⁻

20 1.83×10^{23} C atoms

21 171 g (integer value because no calculator)

22 10.0 g H₂O

23 2.0 mol N₂ > 3.0 mol NH₃ > 25.0 mol H₂ > 1.0 mol N₂H₄

- 24 (a) CH (b) CH₂O
 (c) C₂H₂O₂ (d) C₂H₄
 (e) C₂H₂ (f) CH₂O

25 Na₂S₂O₅

26 CoSO₄·7H₂O

27 C₁₁H₁₆N

28 NH₃

29 6.94 L

30 CdS

31 empirical formula CH; molecular formula C₇H₈

32 empirical formula H₂PO₃; molecular formula H₄P₂O₆

33 C₁₂H₁₈N₂P₂O₁₂ for both empirical and molecular formulas

34 C₂H₃O

35 Let y = mass of chalk in grams.

$$\text{moles of chalk used} = \frac{\text{mass used}}{M_r(\text{CaCO}_3)} \times$$

$$y \frac{\text{g}}{100.09 \text{ g mol}^{-1}}$$

This is the same as the number of moles of carbon atoms used.

$$\text{Therefore the number of carbon atoms used} = \text{moles of chalk} \times (6.02 \times 10^{23} \text{ mol}^{-1}) = \frac{6.02 \times 10^{23} y}{100.09}$$

- 36 (a) 2.50 mol (b) 5.63 mol
 (c) 665.5 g

37 (a) $2\text{C}_2\text{H}_6 + 13\text{O}_2 \rightarrow 8\text{CO}_2 + 10\text{H}_2\text{O}$
 (b) 1.59 g

38 4.355 kg

- 39 (a) $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$
 (b) 92.8%

(c) CaCO₃ is the only source of CO₂; all the CaCO₃ undergoes complete decomposition; all CO₂ released is captured; heating does not cause any change in the mass of the other minerals present.

- 40 (a) 65.2 g (b) 1.3 g H₂

41 5.23 g C₂H₅Cl₂

42 253.2 g theoretical CaSO₄; 77.9%

43 3.16 g ester

44 107 g of C₂H₆ needed

- 45 (a) 2.40 mol (b) 0.0110 mol
 (c) 44 mol

- 46 (a) 35.65 dm³ (b) 5.7 dm³

47 0.652 dm³

48 0.138 mol Br₂ and 0.156 mol Cl₂, so more molecules of Cl₂

49 0.113 dm³

50 0.28 dm³

51 90 kPa

52 16 °C

53 3.0 dm³

54 2.8 dm³

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Table of Contents Chapter 1pearson Chemistry Workbook Answers

1. Understanding the eBook Chapter 1pearson Chemistry Workbook Answers
 - The Rise of Digital Reading Chapter 1pearson Chemistry Workbook Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 1pearson Chemistry Workbook Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chapter 1pearson Chemistry Workbook Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 1pearson Chemistry Workbook Answers
 - Personalized Recommendations
 - Chapter 1pearson Chemistry Workbook Answers User Reviews and Ratings
 - Chapter 1pearson Chemistry Workbook Answers and Bestseller Lists
5. Accessing Chapter 1pearson Chemistry Workbook Answers Free and Paid eBooks

- Chapter 1 Pearson Chemistry Workbook Answers Public Domain eBooks
- Chapter 1 Pearson Chemistry Workbook Answers eBook Subscription Services
- Chapter 1 Pearson Chemistry Workbook Answers Budget-Friendly Options
- 6. Navigating Chapter 1 Pearson Chemistry Workbook Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Chapter 1 Pearson Chemistry Workbook Answers Compatibility with Devices
 - Chapter 1 Pearson Chemistry Workbook Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chapter 1 Pearson Chemistry Workbook Answers
 - Highlighting and Note-Taking Chapter 1 Pearson Chemistry Workbook Answers
 - Interactive Elements Chapter 1 Pearson Chemistry Workbook Answers
- 8. Staying Engaged with Chapter 1 Pearson Chemistry Workbook Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chapter 1 Pearson Chemistry Workbook Answers
- 9. Balancing eBooks and Physical Books Chapter 1 Pearson Chemistry Workbook Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chapter 1 Pearson Chemistry Workbook Answers
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 1 Pearson Chemistry Workbook Answers
 - Setting Reading Goals Chapter 1 Pearson Chemistry Workbook Answers
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 1 Pearson Chemistry Workbook Answers
 - Fact-Checking eBook Content of Chapter 1 Pearson Chemistry Workbook Answers
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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