

Assessment

Chapter Test B - For Practice!

0

Chapter: Atoms: The Building Blocks of Matter

PART I On the line at the left of each statement, write the letter of the choice that best completes the statement or best answers the question.

C

1. The behavior of cathode rays in a glass tube containing gas at low pressure led scientists to conclude that the rays are composed of
- energy.
 - positive particles.
 - ☒ negative particles.
 - neutral particles.

C

2. The basic principles of atomic theory were first conceived by
- Avogadro.
 - Bohr.
 - ☒ Dalton.
 - Rutherford.

C

3. The law of multiple proportions is demonstrated by the compounds
- FeCl_2 and $\text{Fe}(\text{SO}_4)_2$.
 - O_2 and O_3 .
 - ☒ CO and CO_2 .
 - FeCl_2 and $\text{Fe}(\text{NO}_3)_2$.

B

4. Neutral atoms of the same element can differ in
- chemical properties.
 - ☒ mass number.
 - atomic number.
 - number of protons and electrons.

D

5. Dalton's atomic theory helped to explain the law of conservation of mass because it stated that atoms
- cannot combine.
 - are indivisible.
 - all have the same mass.
 - ☒ cannot be created or destroyed.

C

6. Millikan's experiments led to the determination
- that the electron carries no charge.
 - of the electron's positive charge.
 - ☒ of the approximate value of the electron's mass.
 - that the electron has no mass.

Chapter 1skills Practice Answers

Richard H. Loftin



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